

CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

Covalent Bonding

Section 11 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond	valence	sigma bond	valence	p bond
When sharing of electrons occurs, the electrons between atoms that results is called	(a) covalent bond	When each of the atoms is bonded to other electrons	(b) valence	When two or more atoms bond by means of electrons sharing the resulting particle is called an

(c) **molecule** If the electrons shared are neutral between the two atoms, the molecule is called an (d) **sigma bond** If the sharing involves the overlap of parallel orbitals, the molecule is called an (e) **p bond**

In your textbook, read about ionic and covalent bonds and bond strength.

Circle the letter of the choice that best completes the statement or answers the question.

(f) In which two elements is it a hydrogen, nitrogen, and oxygen atom(s) are?

- an ionic bond
- a molecule containing two atoms

(g) The energy electrons are shared in a double covalent bond?

- one
- two
- four
- five

(h) Bond length is the distance between

- two molecules of the same substance
- the electrons in two different atoms

(i) What is the following relationship referring to bond length is generally correct?

- the shorter the bond, the stronger the bond
- the shorter the bond, the weaker the bond
- the shorter the bond, the stronger the bond
- the shorter the bond, the lower the bond strength is

CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

Section 12 Naming Molecules

In your textbook, read about how binary compounds and salts are named from their formulas.

For each statement below, write true or false.

(a) _____ 1. Binary molecular compounds are generally composed of a metal and a nonmetal.

(b) _____ 2. The general structure of a binary compound is shown as _____ the symbol for the metal is _____ to the left of the symbol for the nonmetal.

(c) _____ 3. The prefix *mono* is used when there are two atoms.

(d) _____ 4. The prefix *di* is used when there are two atoms.

(e) _____ 5. In naming the first element in a formula, the prefix *mono* is not used.

(f) _____ 6. For binary acids, the hydrogen part of the compound is named using the prefix *hydro*.

(g) _____ 7. An optional prefix is used for elements.

(h) _____ 8. If the name of the element is an optional prefix in only the acid name, it comes after the prefix *hydro*.

In your textbook, read about naming organic compounds and isomers.

For each item in Column A, write the letter of the matching item in Column B.

Column A

1. **hydrocarbon**

2. **nitro compound**

3. **alcohol**

4. **ether**

5. **ketone**

6. **aldehyde**

7. **carboxylic acid**

8. **alicyclic compound**

9. **aliphatic acid**

10. **aromatic compound**

11. **alcohol**

12. **ether**

13. **ketone**

14. **aldehyde**

15. **carboxylic acid**

16. **alicyclic compound**

17. **aromatic compound**

18. **aliphatic acid**

Name _____ Date _____	Name _____ Date _____
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Column A

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(g) _____ 7. **carboxylic acid**

(h) _____ 8. **alicyclic compound**

(i) _____ 9. **aliphatic acid**

(j) _____ 10. **aromatic compound**

(k) _____ 11. **alcohol**

(l) _____ 12. **ether**

(m) _____ 13. **ketone**

(n) _____ 14. **aldehyde**

(o) _____ 15. **carboxylic acid**

(p) _____ 16. **alicyclic compound**

(q) _____ 17. **aromatic compound**

(r) _____ 18. **aliphatic acid**

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J Dewey

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