

Name _____ Date _____ Class _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27 in the space below the table.



The molar masses of O_2 , NO , and N_2 are 32.00 g/mol, 30.01 g/mol, and 28.02 g/mol, respectively.

Amount of O_2	Amount of N_2	Amount of NO	Limiting Reactant	Amount and State of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
4 molecules	4 molecules	4 molecules	N_2	2 molecules O_2
2 molecules	8 molecules	4 molecules	2. O_2	3. 4 molecules NO
1.00 mol	2.00 mol	4. 2.00 mol	5. none	6. none
4.00 mol	4.00 mol	7. 4.00 mol	8. NO	9. 2.00 mol O_2
1.00 mol	7.00 mol	10. 7.00 mol	11. NO	12. 1.50 mol O_2
1.00 mol	4.00 mol	13. 2.00 mol	14. O_2	15. 2.00 mol NO
0.500 mol	0.200 mol	16. 0.200 mol	17. NO	18. 0.400 mol O_2
32.00 g	80.02 g	19. 92.02 g	20. none	21. none
16.00 g	80.02 g	22. 46.01 g	23. O_2	24. 58.12 g NO
16.00 g	28.02 g	25. 28.76 g	26. O_2	27. 1.24 g NO

balanced equation mole ratio = 2 mol NO /1 mol O_2

$16.00 \text{ g } \text{O}_2 \times 1 \text{ mol } \text{O}_2 / 32.00 \text{ g } \text{O}_2 = 0.5125 \text{ mol } \text{O}_2$

$28.00 \text{ g } \text{N}_2 \times 1 \text{ mol } \text{N}_2 / 28.01 \text{ g } \text{N}_2 = 0.9999 \text{ mol } \text{N}_2$

actual mole ratio = $0.9999 \text{ mol } \text{NO} / 0.5125 \text{ mol } \text{O}_2 = 2.132 \text{ mol } \text{NO} / 1.000 \text{ mol } \text{O}_2$

Because the actual mole ratio of $\text{NO}:\text{O}_2$ is larger than the balanced equation mole ratio of $\text{NO}:\text{O}_2$, there is an excess of NO ; O_2 is the limiting reactant.

Mass of NO used = $0.5125 \text{ mol } \text{O}_2 \times 2 \text{ mol } \text{NO} / 1 \text{ mol } \text{O}_2 = 0.4250 \text{ mol } \text{NO}$

$0.4250 \text{ mol } \text{NO} \times 30.01 \text{ g } \text{NO} / 1 \text{ mol } \text{NO} = 12.76 \text{ g } \text{NO}$

Mass of NO_2 produced = $0.4250 \text{ mol } \text{NO}_2 \times 46.01 \text{ g } \text{NO}_2 / 1 \text{ mol } \text{NO}_2 = 19.56 \text{ g } \text{NO}_2$

Excess $\text{NO} = 28.00 \text{ g } \text{NO} - 12.76 \text{ g } \text{NO} = 15.24 \text{ g } \text{NO}$

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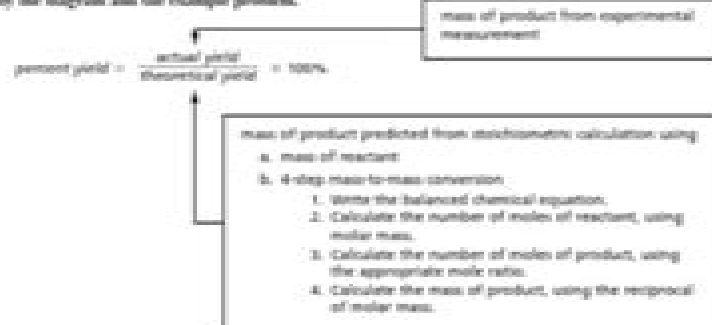
CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.4 Percent Yield

In your textbook, read about the yield of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of gallium oxide, a substance used in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the oxide from a 7.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 187.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield: unknown 7.42 g Ga_2O_3
- mass of reactant: 7.00 g Ga correct
- number of moles of reactant: $7.00 \text{ g } \text{Ga} \times \frac{69.72 \text{ g } \text{Ga}}{1 \text{ mol } \text{Ga}} = 7.00 \text{ g } \text{Ga} \times 1 \text{ mol } \text{Ga} / 69.72 \text{ g } \text{Ga}$
- number of moles of product: $0.100 \text{ mol } \text{Ga} \times \frac{2 \text{ mol } \text{Ga}_2\text{O}_3}{1 \text{ mol } \text{Ga}} = 0.100 \text{ mol } \text{Ga} \times \frac{2 \text{ mol } \text{Ga}_2\text{O}_3}{1 \text{ mol } \text{Ga}}$
- theoretical yield: $0.100 \text{ mol } \text{Ga}_2\text{O}_3 \times \frac{187.44 \text{ g } \text{Ga}_2\text{O}_3}{1 \text{ mol } \text{Ga}_2\text{O}_3} = 18.74 \text{ g } \text{Ga}_2\text{O}_3$ correct
- percent yield: $\frac{7.42 \text{ g } \text{Ga}_2\text{O}_3}{18.74 \text{ g } \text{Ga}_2\text{O}_3} \times 100 = 7.42 \text{ g } \text{Ga}_2\text{O}_3 / 18.74 \text{ g } \text{Ga}_2\text{O}_3 \times 100$

Study Guide For Content Mastery The Mole

R Sanford



Study Guide For Content Mastery The Mole:

Connecting Self-regulated Learning and Performance with Instruction Across High School Content Areas

Maria K. DiBenedetto, 2018-07-23 This book shows how principles of self regulated learning are being implemented in secondary classrooms The 14 chapters are theoretically driven and supported by empirical research and address all common high school content areas The book comprises 29 lesson plans in English language arts natural and physical sciences social studies mathematics foreign language art music health and physical education Additionally the chapters address students with special needs technology and homework Each chapter begins with one or more lesson plans written by master teachers followed by narratives explaining how the lesson plans were implemented The chapters conclude with an analysis written by expert researchers of the self regulated learning elements in the lessons Each lesson and each analysis incorporate relevant educational standards for that area Different types of high schools in several states serve as venues This powerful new book edited by Maria K DiBenedetto provides a unique and invaluable resource for both secondary teachers and researchers committed to supporting adolescents in the development of academic self regulation Each chapter is jointly written by teachers who provide a wealth of materials including lesson plans and researchers who situate these lesson plans and academic self regulation goals within the larger work on self regulation The topics covered are far broader than any other book I have seen in terms of developing academic self regulation covering over a dozen content areas including literacy mathematics social studies the sciences and the arts Teachers and scholars alike will find this book a must read Karen Harris EdD Arizona State University A practical and magnificent blend of educational research and application This book goes beyond presenting the findings of research on self regulation by connecting detailed strategies that align with the standards to the research DiBenedetto et al clearly illustrate how to develop self regulated learners in the classroom A refreshing must read for all secondary educators and educational researchers seeking to be well grounded in education research and practical application techniques Heather Brookman PhD Fusion Academy Park Avenue Self regulated learning is a research based process by which teachers help students realize their own role in the learning process Connecting Self Regulated Learning and Performance with Instruction Across High School Content Areas consists of model teachers lessons and analyses by prominent educational psychologists in the field of self regulated learning The book provides teachers with the tools needed to increase students awareness of learning and inspires all educators to use self regulated learning to promote engagement motivation and achievement in their students The book also provides administrators with the principles needed to infuse evidenced based self regulated learning into their curriculum and instruction I highly recommend the book Marty Richburg Northside High School [Encyclopedia of Giftedness, Creativity, and Talent](#) Barbara Kerr, 2009-06-15 Library Journal Best Reference 2009 Educators have begun to recognize that identifying and accommodating students who are slower to learn may leave behind those who learn rapidly and think creatively The nature of intelligence the sources of

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