

**Ch 5 Ions and Ionic Compounds Study Guide****Multiple Choice***Identify the choice that best completes the statement or answers the question.*

1. Which of the following electron configurations belongs to an element that is NOT chemically reactive?  
a.  $1s^2 2s^2 2p^2$       c.  $1s^2 2s^2 2p^6$   
b.  $1s^2 2s^2 2p^6 3s^2 3p^6$       d.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$

2. In the compound sodium fluoride, NaF, the sodium atom loses one electron and the fluorine atom gains one electron to form ions that have electron configurations similar to  
a. helium.      c. neon.  
b. oxygen.      d. calcium.

3. In many compounds, atoms of main-group elements form ions so that the number of electrons in the outermost energy levels of each ion is  
a. 2.      c. 8.  
b. 6.      d. 10.

4. The electron configuration of nitrogen is  $1s^2 2s^2 2p^3$ . How many more electrons does nitrogen need to have an electron configuration similar to neon?  
a. 1      c. 5  
b. 3      d. 8

5. Atoms of copper and iron  
a. generally form stable bonds with transition elements.  
b. have stable electron configurations.  
c. tend to form cations.  
d. tend to form anions.

6. An anion  
a. is an ion with a negative charge.  
b. attracts ions with negative charges.  
c. results when an alkaline-earth metal loses one of its two outermost electrons.  
d. has more protons than electrons.

7. When the octet rule is satisfied, the outermost \_\_\_\_\_ are filled.  
a. *d* and *f* orbitals      c. *s* and *d* orbitals  
b. *s* and *p* orbitals      d. *d* and *p* orbitals

8. The elements of the \_\_\_\_\_ group satisfy the octet rule without forming compounds.  
a. halide      c. alkali metal  
b. noble gas      d. alkaline-earth metal

9. Once an atom has full *s* and *p* orbitals in its outermost energy level,  
a. it is highly reactive only with alkali metals.  
b. it is highly reactive only with halogens.  
c. it can be combined with most elements.  
d. it has a stable octet and is unreactive.

# Study Guide For Ions And Ionic Compounds

**Susan McMurry, John McMurry**

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## **Table of Contents Study Guide For Ions And Ionic Compounds**

1. Understanding the eBook Study Guide For Ions And Ionic Compounds
  - The Rise of Digital Reading Study Guide For Ions And Ionic Compounds
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Ions And Ionic Compounds
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide For Ions And Ionic Compounds
  - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide For Ions And Ionic Compounds
  - Personalized Recommendations
  - Study Guide For Ions And Ionic Compounds User Reviews and Ratings

- Study Guide For Ions And Ionic Compounds and Bestseller Lists
- 5. Accessing Study Guide For Ions And Ionic Compounds Free and Paid eBooks
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  - Study Guide For Ions And Ionic Compounds eBook Subscription Services
  - Study Guide For Ions And Ionic Compounds Budget-Friendly Options
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  - ePUB, PDF, MOBI, and More
  - Study Guide For Ions And Ionic Compounds Compatibility with Devices
  - Study Guide For Ions And Ionic Compounds Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Study Guide For Ions And Ionic Compounds
  - Highlighting and Note-Taking Study Guide For Ions And Ionic Compounds
  - Interactive Elements Study Guide For Ions And Ionic Compounds
- 8. Staying Engaged with Study Guide For Ions And Ionic Compounds
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Study Guide For Ions And Ionic Compounds
- 9. Balancing eBooks and Physical Books Study Guide For Ions And Ionic Compounds
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Study Guide For Ions And Ionic Compounds
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide For Ions And Ionic Compounds
  - Setting Reading Goals Study Guide For Ions And Ionic Compounds
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide For Ions And Ionic Compounds
  - Fact-Checking eBook Content of Study Guide For Ions And Ionic Compounds
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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