

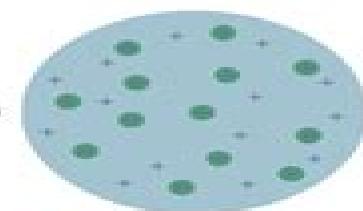
## Atom History

- John Dalton was the first scientist to make a model of the atom. He said that two types of atoms form a compound. He made a few postulates and formed the law of conservation of matter.
- Law of conservation of matter - Matter cannot be created or destroyed. It can only be changed.
- Postulates
  1. All matter is composed of atoms.
  2. Atoms are hollow spheres that cannot be broken down.
  3. Atoms of different elements can combine to form a compound.
  4. Atoms of the same element are always identical.

Dalton's Atomic Theory



- J.J. Thomson was the second person to approach the atom problem.
- J.J. Thomson used the cathode ray tube experiment.
- Cathode ray tube - a high voltage vacuum with electric field
- In Thomson's experiment, the cathode ray bent towards the positive end which suggested that something attracted and there was something inside the atom.
- The ray tube discovered the thing that Dalton missed which was charges.
- The experiment proved that there are electrons.
- Thomson's model is aka the plum pudding model.

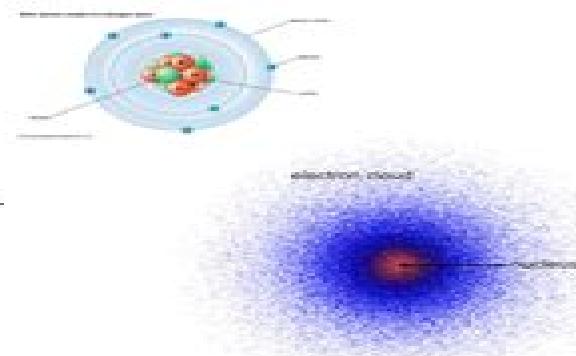


- Ernst Rutherford was the third person to make a change to the atom.
- He conducted the gold foil experiment.
- Throughout the experiment, he observed that most alpha particles pass through the gold foil and that some reflected and bounced back.
- This experiment proved that an atom is mostly empty space because most of the particles passed through. However, the atom DOES have a positive nucleus because some of the particles reflected / deflected back.
- His model was the nuclear model.
- The problem with the model is that positive and negative charges charge so this would mean that the electrons would track into the nucleus.
- ESPN - mostly empty space and positive nucleus



- Bohr was the third person to change the atom model.
- He added that electrons exist in orbits or shells.
- He said that each orbit has a specific amount of energy.
- Most used model today.

- Shrodinger was the fourth person to change the atom model.
- He said that electrons exist in orbitals around the nucleus.
- Orbitals - areas where it is most likely to find electrons.



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