

Atomic Structure and the Periodic Table Study Guide

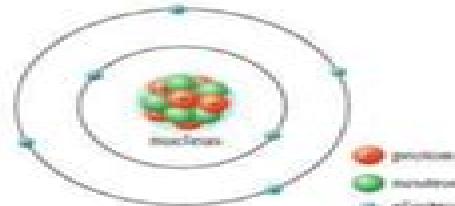
## 1. What is an atom?

Smallest particle of an element that retains the properties of that element.

## 2. Complete the following chart about subatomic particles.

Name	Charge	Symbol	Location	Relative Mass
Proton	+	$p^+$	Nucleus	1 amu
Electron	-	$e^-$	Electron cloud	0
neutron	0	$n^0$	nucleus	1 amu

## 3. Draw an atom and label the subatomic particles.

4. Which subatomic particle is used to determine the identity of an element?  
protons5. Where is most of the mass of the atom? What particle(s) are in this area?  
Nucleus; protons and neutrons

## 6. What is an atom called after it has gained or lost electrons? \_\_\_\_\_ion\_\_\_\_\_

## 7. What do you call an atom if it has more or less neutrons than normal? \_\_\_\_\_isotope\_\_\_\_\_

## 8. Label the information provided in the periodic table.

8	→	Atomic number
O	→	Symbol
Oxygen	→	Name
15.999	→	Atomic mass

## 9. From a box on the periodic table, how do you determine the number of protons, electrons, and neutrons?

1) #  $p^+$  = \_\_\_\_\_ atomic number \_\_\_\_\_2) #  $e^-$  = \_\_\_\_\_ number of protons \_\_\_\_\_3) #  $n^0$  = \_\_\_\_\_ atomic mass - atomic number \_\_\_\_\_

## 10. What does the atomic number represent? # protons in the nucleus

## 11. What does the atomic mass represent? # particles in the nucleus (protons &amp; neutrons)

# Study Guide Atoms And Elements

**Susan McMurry, John McMurry**

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