

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

solvent	liquid	solvent
insoluble	solid	solvent

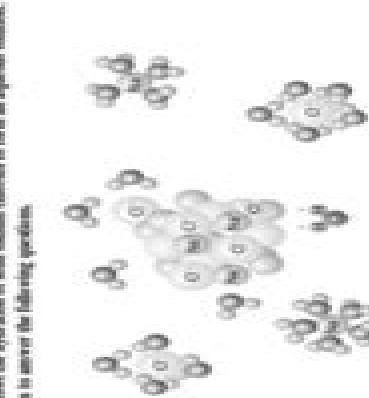
After reading (1) _____ solution _____ of solute is dissolved in solvent.

(2) _____ solute _____ is the (3) _____ solute _____ and solute is the

(4) _____ solution. A solution that does not dissolve is
 (5) _____ insoluble. (6) _____ liquid _____ solution is the most common
 type of solution. If one liquid is soluble in another liquid, such as water and oil, water
 can dissolve in (7) _____ insoluble _____ because if one liquid is insoluble in another
 the liquid is an (8) _____ insoluble.

Read about solubility in aqueous solution in your textbook.

Underline the questions to answer the following questions.



1. Indicate in solute is when the solute is in a solution. What is solute?

Solute is the process of surrounding solute particles with solvent particles to form a solution.

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Section 15.1 content**STUDY GUIDE FOR CONTENT MASTERY**

16. An sodium chloride dissolved in water. What happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

17. Explain the attraction of the water molecules around the sodium ions and chloride ions.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atom). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.

19. In a glass cup, add the rest of solute may be dissolved.

Starting at the bottom of the solution, breaking the solution into smaller pieces, and heating the solution.

In your textbook, read about heat of solution, solubility and factors that affect solubility for each solute and solvents, with tips or jobs.

True _____ 16. The overall energy change that occurs when a solute dissolves is called the heat of solution.

False _____ 17. Solubility is a property of the solute, amount of either the solute or a given amount of solute at a specified temperature and pressure.

False _____ 18. Solubility continues to drop as the solution has a high heat of solution.

True _____ 19. In a saturated solution, solute and solvents are in equilibrium.

True _____ 20. Additional solute can be dissolved in an unsaturated solution.

True _____ 21. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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**William B. Walstad, Campbell R.
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