

Properties of water

Study Guide with Summary questions

Water is a polar molecule because the hydrogen atoms are more positive than the oxygen atom, and the oxygen atom is more negative than the hydrogen atoms.

Hydrogen Bonding Water molecules are able to form hydrogen bonds with other water molecules. This is because the oxygen atom in one water molecule has a partial negative charge, and the hydrogen atoms in another water molecule have partial positive charges.

Dielectric Constant Dielectric constant is the ratio of the charge density of a material to the charge density of a vacuum. Water has a dielectric constant of 80, which is higher than most other materials.

Surface Tension Water has a high surface tension because the hydrogen bonds between water molecules are very strong.

Heat Capacity Water has a high specific heat capacity because it has a high heat capacity.

Viscosity Water has a low viscosity because it has a low density.

Adhesion and Cohesion Water molecules are attracted to other polar molecules. This is because water molecules have partial positive charges.

Water is a polar molecule, so it has a partial positive charge on the oxygen atom and partial positive charges on the hydrogen atoms.

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Properties of Water



Universal Solvent

Water is often referred to as the "universal solvent" due to its ability to dissolve many substances. This property is crucial for life on Earth, allowing for the breakdown of complex molecules into smaller, more manageable components.

Water is a polar molecule, so it has a partial positive charge on the oxygen atom and partial positive charges on the hydrogen atoms. The oxygen atom has a partial negative charge, and the hydrogen atoms have partial positive charges. This property allows water to dissolve polar substances by forming hydrogen bonds with the polar molecules.

Water's polarity is exemplified in its ability to dissolve ionic compounds. The ionic bonds between the cations and anions are broken, allowing the water molecules to surround the ions and stabilize them.

Water's polarity is also exemplified in its ability to dissolve polar molecules. The partial positive charge on the oxygen atom and partial negative charges on the hydrogen atoms allow water molecules to surround the polar molecules and stabilize them.

Hydrogen

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Hydrogen

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