

Unit 1 Chemistry Study Guide

History of the Atom:

Democritus (450 BCE):

- Matter is not infinitely divisible.
- There are single particles that cannot be divided further, called **atomos**.

Aristotle (380 BCE):

- Matter is infinitely divisible.
- Matter is called **hyle**.

Antoine Lavoisier (1774):

- **Law of Conservation of Mass:** Mass cannot be created or destroyed.

Joseph Proust (1779):

- **Law of Definite Proportions:** All samples of a pure compound contain the same elements in the same proportion by mass (no matter quantity).

John Dalton (1803 – 1808):

- Observed work of Lavoisier and Proust
- Led Dalton to develop the first **atomic theory**.
 - o **Daltons Atomic Theory:**
 - o All matter is composed of extremely small particles called atoms.
 - o All atoms of a given element are identical, having the same size, mass, and chemical properties. The atoms of one element are different from the atoms of all other elements.
 - o Atoms cannot be subdivided, created, or destroyed.
 - o Atoms to different elements can combine in simple whole number ratios to form chemical compounds.
 - o In chemical reactions, atoms are combined, separated or rearranged.
 - o **Law of Multiple Proportions:** Elements can combine in different proportions. However, the different compounds will have different properties but the ratio of atoms is always a whole number.

Billiard Ball Model:



J.J. Thompson (1906):

- **Cathode Ray Experiment –**
 - o Part 1: Thompson ran electricity through a cathode ray tube and when he put the negatively charged end of a magnet near the tube, he saw the ray of electricity

Unit 1 Chemistry Study Guide

Karin Nielsen-Saines

Unit 1 Chemistry Study Guide:

Ccea as Chemistry Student Unit Guide Alyn McFarland,2012 Written by a senior examiner Alyn G McFarland this CCEA AS Chemistry Student Unit Guide is the essential study companion for Unit 1 Basic Concepts in Physical and Inorganic Chemistry This full colour book includes all you need to know to prepare for your unit exam aaaa clear guidance on the content of the unit with topic summaries knowledge check questions and a quick reference index aaaa examiner's advice throughout so you will know what to expect in the exam and will be able to demonstrate the skills required aaaa exam style questions with graded student responses so you can see clearly what is required to get a better grade

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for McMurry's Organic Chemistry, Fifth Edition Susan McMurry,2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Prize winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines

Study Guide General Chemistry, Second Edition,
Becker/Wentworth Becker,Benjamin L. Carroll,1980 **PSI Study Guide for Organic Chemistry** by Douglas C. Neckers,

Michael P. Doyle E. C. Blossey,1977 **Edexcel AS Chemistry Student Unit Guide New Edition: Unit 1 The Core**

Principles of Chemistry Rod Beavon,George Facer,2012-05-18 Written by a senior examiner Rod Beavon and revised by George Facer this Edexcel AS Chemistry Student Unit Guide is the essential study companion for Unit 1 The Core Principles of Chemistry This full colour book includes all you need to know to prepare for your unit exam Clear guidance on the content of the unit with topic summaries knowledge check questions and a quick reference index Examiner's advice throughout so you will know what to expect in the exam and will be able to demonstrate the skills required Exam style questions with graded student responses so you can see clearly what is required to get a better grade

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