

Unit 1 Chemistry Study Guide

History of the Atom:

Democritus (450 BCE):

- Matter is not infinitely divisible.
- There are single particles that cannot be divided further, called **atomos**.

Aristotle (380 BCE):

- Matter is infinitely divisible.
- Matter is called **hyle**.

Antoine Lavoisier (1774):

- **Law of Conservation of Mass:** Mass cannot be created or destroyed.

Joseph Proust (1779):

- **Law of Definite Proportions:** All samples of a pure compound contain the same elements in the same proportion by mass (no matter quantity).

John Dalton (1803 – 1808):

- Observed work of Lavoisier and Proust
- Led Dalton to develop the first **atomic theory**.
- **Dalton's Atomic Theory:**
 - o All matter is composed of extremely small particles called atoms.
 - o All atoms of a given element are identical, having the same size, mass, and chemical properties. The atoms of one element are different from the atoms of all other elements.
 - o Atoms cannot be subdivided, created, or destroyed.
 - o Atoms of different elements can combine in simple whole number ratios to form chemical compounds.
 - o In chemical reactions, atoms are combined, separated or rearranged.
- **Law of Multiple Proportions:** Elements can combine in different proportions. However, the different compounds will have different properties but the ratio of atoms is always a whole number.

Billiard Ball Model:



atom

J.J. Thompson (1906):

- **Cathode Ray Experiment –**
 - o Part 1: Thompson ran electricity through a cathode ray tube and when he put the negatively charged end of a magnet near the tube, he saw the ray of electricity

Unit 1 Chemistry Study Guide Answers

LP Steffe



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Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition Susan McMurry, 2000

Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Prize winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines

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Chemistry Study Guide Sixth Edition Darrell D. Ebbing, 1999 **Student Study Guide for Biology [by]**
Campbell/Reece/Mitchell Martha R. Taylor, 1999 Prentice Hall Chemistry ,2000 **Abstracts of Theses, Researches**
in Progress, and Bibliography of Publications ,1930 **The Student in Pharmacy** Eugene Gustave Eberle, 1913

Open Science Shelagh Ross, Eileen Scanlon, 1995-11-28 This book explores the particular challenges of teaching and learning science at a distance It discusses the principles of instructional design informed by studies of the ways in which conceptual and process based skills are developed by science learners and suggests how structural issues in the selection of course content may be addressed Focusing mainly on examples drawn from undergraduate teaching successive chapters illustrate a wide range of instructional media and techniques and show how multiple media courses can be assembled to form a coherent whole The issues of student assessment course evaluation and quality audit are also examined **Resources in**

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