

Unit 3 Worksheet 3 – Quantitative Energy Problems

Energy constants (H₂O)

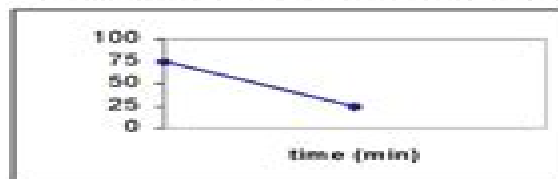
334 J/g	Heat of fusion (melting or freezing) H _f
2260 J/g	Heat of vaporization (evaporating or condensing) H _v
2.1 J/g°C	Heat capacity (c) of solid water
4.18 J/g°C	Heat capacity (c) of liquid water

$$Q = m \cdot c \cdot \Delta T$$

$$Q = m \cdot H_v \text{ or } m \cdot H_f$$

For each of the problems sketch a warming or cooling curve to help you decide which equation(s) to use to solve the problem. Keep a reasonable number of sig figs in your answers.

1. A cup of coffee (140 g) cools from 75°C down to comfortable room temperature 20.°C. How much energy does it release to the surroundings?



$$Q = 140g \cdot 4.18 \frac{J}{g^{\circ}C} \cdot 55^{\circ}C = 32,000J \text{ or } 32 \text{ kJ}$$

2. Suppose during volleyball practice, you lost 2.0 lbs of water due to sweating. If all of this water evaporated, how much energy did the water absorb from your body? Express your answer in kJ. 2.2 lbs = 1.0 kg



$$2.0lbs \cdot \frac{1kg}{2.2lbs} \cdot \frac{1000g}{1kg} = 910g$$

$$Q = 910g \cdot 2260 \frac{J}{g} = 2,100,000J \text{ or } 2100 \text{ kJ}$$

3. Suppose that during the Icy Hot lab that 65 kJ of energy were transferred to 450 g of water at 20°C. What would have been the final temperature of the water?



$$Q = m \cdot c \cdot \Delta T(^{\circ}C)$$

$$\Delta T = \frac{Q}{m \cdot c} = \frac{65,000J}{450g \cdot 4.18 \frac{J}{g^{\circ}C}} = 34.6^{\circ}C = 35^{\circ}C$$

Unit 3 Worksheet 4 Quantitative Energy Problems

Guillaume Favre



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