

HONORS - VSEPR Theory - Molecular Geometry Reference Sheet**- Unshared pair - Memorize**

- In electron dot diagrams for covalent compounds, remember that there are two types of electron pairs: shared and unshared.

Electron Domains

- Shared pairs of electrons are **BONDS** and are shown as a connection between the two atoms bonding.
- Unshared pairs of electrons stay on the atom they originate from and are **NOT bonded**.
- Electron dot diagrams are 3-dimensional so remember that more than 2 electron domains require you to move forward or backwards in space with your drawing.

*Central atom is usually the least electronegative element in the formula (except Hydrogen) and/or the element with the most open bond locations.

Names of types of chemical bonds used to draw 3D molecular structures:

Orientations of chemical bonds indicated by these styles:

normal bond _____

bond lies in the plane of the paper (or screen, when viewed electronically)

dashed bond _____

bond extends backwards, away from the viewer, so effectively "out of" the paper (or screen)

wedged bond _____

bond extends forwards, towards the viewer, so effectively "out of" the paper (or screen)

* Unshared pairs repel w/ 2-3° more than bonds

Name of Shape	Molecular Shape	# of Total Electron Domains	# of atoms bonded to central atom	Unshared pairs on central atom	Bond Angles	Example
Linear		2	1 or 2	0	180°	CO_2
Polyatomic Trigonal-planar + Boron		3	3	0	120°	BF_3
Bent (V-Shaped)		3	2	1	117°	SO_2
		4	2	2	105°	H_2O
Tetrahedral * 3D		4	4	0	109.5°	CH_4
Trigonal-pyramidal * 3D		4	3	1	107°	NH_3

Unit 6 Molecular Clock Study Guide Answers

St. James's Theatre (London, England)

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- Chapter 2: Essential Elements of Unit 6 Molecular Clock Study Guide Answers
- Chapter 3: Unit 6 Molecular Clock Study Guide Answers in Everyday Life
- Chapter 4: Unit 6 Molecular Clock Study Guide Answers in Specific Contexts
- Chapter 5: Conclusion

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