

HONORS - VSEPR Theory - Molecular Geometry Reference Sheet

Unshared pair

Memorize

Electron Domains

- In electron dot diagrams for covalent compounds, remember that there are two types of electron pairs: shared and unshared.

- Shared pairs of electrons are **BONDS** and are shown as a connection between the two atoms bonding.
- Unshared pairs of electrons stay on the atom they originate from and are **NOT** bonded.

- Electron dot diagrams are 3-dimensional so remember that more than 2 electron domains require you to move forward or backwards in space with your drawing.

Names of types of chemical bonds used to draw 3D molecular structures:

"normal" bond

dashed bond

wedged bond

Orientations of chemical bonds indicated by these styles:

bond lies in the plane of the paper (or screen, when viewed electronically)

bond extends backwards, away from the viewer, so effectively "into" the paper (or screen)

bond protrudes forwards, towards the viewer, so effectively "out of" the paper (or screen)

*** Unshared pairs repel ~ 2-3° more than bonds**

*Central atom is usually the least electronegative element in the formula (except Hydrogen) and/or the element with the most open bond locations.

Name of Shape	Molecular Shape	# of Total Electron Domains	# of atoms bonded to central atom	Unshared pairs on central atom	Bond Angles	Example
Linear		2	1 or 2 $\text{O}=\text{O}$	0	180°	CO_2
Trigonal planar + Boron		3	3	0	120°	BF_3
Bent (V-Shaped)		3	2	1	117°	SO_2
		4	2	2	105°	H_2O
Tetrahedral * 3D		4	4	0	109.5°	CH_4
Trigonal pyramidal * 3D		4	3	1	107°	NH_3

Breaks Octet

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- Chapter 5: Conclusion

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