

Chemical Bonding

SECTION 3

SHORT ANSWER Answer the following questions in the space provided.

- A** The notation for sodium chloride, NaCl, stands for one
 (a) formula unit. (c) crystal.
 (b) molecule. (d) atom.
- D** In a crystal of an ionic compound, each cation is surrounded by a number of
 (a) molecules. (c) dipoles.
 (b) positive ions. (d) negative ions.
- B** Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is
 (a) higher in potential energy. (c) equal in potential energy.
 (b) lower in potential energy. (d) unstable.
- B** The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?
 (a) Compound A is not an ionic compound.
 (b) It will be more difficult to break the bonds in compound A than those in compound B.
 (c) Compound B has larger crystals than compound A.
 (d) Compound A has larger crystals than compound B.
- B** The forces of attraction between molecules in a molecular compound are generally
 (a) stronger than the attractive forces among formula units in ionic bonding.
 (b) weaker than the attractive forces among formula units in ionic bonding.
 (c) approximately equal to the attractive forces among formula units in ionic bonding.
 (d) equal to zero.
- Describe the force that holds two ions together in an ionic bond.

Very strong overall attraction between positive and negative charges

- What type of energy best represents the strength of an ionic bond?

Lattice energy

- What types of bonds are present in an ionic compound that contains a polyatomic ion?

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Table of Contents Unit 6 Section D Content Guide Chemistry

1. Understanding the eBook Unit 6 Section D Content Guide Chemistry
 - The Rise of Digital Reading Unit 6 Section D Content Guide Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 6 Section D Content Guide Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 6 Section D Content Guide Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 6 Section D Content Guide Chemistry
 - Personalized Recommendations
 - Unit 6 Section D Content Guide Chemistry User Reviews and Ratings
 - Unit 6 Section D Content Guide Chemistry and Bestseller Lists
5. Accessing Unit 6 Section D Content Guide Chemistry Free and Paid eBooks
 - Unit 6 Section D Content Guide Chemistry Public Domain eBooks
 - Unit 6 Section D Content Guide Chemistry eBook Subscription Services

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- Unit 6 Section D Content Guide Chemistry Budget-Friendly Options
6. Navigating Unit 6 Section D Content Guide Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 6 Section D Content Guide Chemistry Compatibility with Devices
 - Unit 6 Section D Content Guide Chemistry Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 6 Section D Content Guide Chemistry
 - Highlighting and Note-Taking Unit 6 Section D Content Guide Chemistry
 - Interactive Elements Unit 6 Section D Content Guide Chemistry
 8. Staying Engaged with Unit 6 Section D Content Guide Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 6 Section D Content Guide Chemistry
 9. Balancing eBooks and Physical Books Unit 6 Section D Content Guide Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 6 Section D Content Guide Chemistry
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Unit 6 Section D Content Guide Chemistry
 - Setting Reading Goals Unit 6 Section D Content Guide Chemistry
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Unit 6 Section D Content Guide Chemistry
 - Fact-Checking eBook Content of Unit 6 Section D Content Guide Chemistry
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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