

## Unit 1 Chemistry Study Guide

### History of the Atom:

#### Democritus (450 BCE):

- Matter is not infinitely divisible.
- There are single particles that cannot be divided further, called **atomos**.

#### Aristotle (380 BCE):

- Matter is infinitely divisible.
- Matter is called **hyle**.

#### Antoine Lavoisier (1774):

- **Law of Conservation of Mass:** Mass cannot be created or destroyed.

#### Joseph Proust (1779):

- **Law of Definite Proportions:** All samples of a pure compound contain the same elements in the same proportion by mass (no matter quantity).

#### John Dalton (1803 – 1808):

- Observed work of Lavoisier and Proust
- Led Dalton to develop the first **atomic theory**.
- **Dalton's Atomic Theory:**
  - o All matter is composed of extremely small particles called atoms.
  - o All atoms of a given element are identical, having the same size, mass, and chemical properties. The atoms of one element are different from the atoms of all other elements.
  - o Atoms cannot be subdivided, created, or destroyed.
  - o Atoms of different elements can combine in simple whole number ratios to form chemical compounds.
  - o In chemical reactions, atoms are combined, separated or rearranged.
- **Law of Multiple Proportions:** Elements can combine in different proportions. However, the different compounds will have different properties but the ratio of atoms is always a whole number.

Billiard Ball Model:



atom

J.J. Thompson (1906):

- **Cathode Ray Experiment –**
  - o Part 1: Thompson ran electricity through a cathode ray tube and when he put the negatively charged end of a magnet near the tube, he saw the ray of electricity

# Unit 7 Study Guide Chemistry Answers

**Herbert Drapkin**



## Unit 7 Study Guide Chemistry Answers:

Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition Susan McMurry, 2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Price winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines

**College Biology I** James Hall Zimmerman, Sophie E. Merritt, 1963 *Student Study Guide and Solutions Manual to Accompany General, Organic, and Biochemistry* Katherine J. Denniston, Denniston Katherine, 2006 Chemical Processes Open University. S334 Course Team, 1978 **The Publishers' Trade List Annual**, 1904 **El-Hi Textbooks in Print**

, 1984 **Gage Canadian GED Study Guide Series** Kathy Austin, 1998 **Teacher's Manual and Resource Guide for Exploring the Sciences** Herbert Drapkin, 1964 *Research in Education*, 1974 Resources in Education, 1985

**Manual for PSI Proctors**, 1975 **CHEMISTRY SILBERBERG**, 2003 **AAAS Science Book List, 1978-1986** Kathryn Wolff, Susan M. O'Connell, Valerie J. Montenegro, 1986 A selected and annotated list of science and mathematics books which supplements the AAAS science book list 3rd ed 1970 and the AAAS science book list supplement 1978

*School*, 1922 **Cooperative Chemistry Lab Manual** Cooper, 2005-02 The laboratory course described in the lab manual emphasizes experimental design data analysis and problem solving Inherent in the design is the emphasis on communication skills both written and oral Students work in groups on open ended projects in which they are given an initial scenario and then asked to investigate a problem There are no formalized instructions and students must plan and carry out their own investigations *Peer-Led Team Learning* David Gosser, Victor S. Strozak, Mark S. Cracolice, 2006 The Workshop Program is a collaboration of teachers learning specialists and students that develops and implements a Peer Led Team Learning model of teaching science Students who have done well in the course are selected and trained for leadership roles These peer leaders guide the work of teams of 6-8 students to solve carefully structured problems in weekly Workshop sessions **Whitaker's Books in Print**, 1998 Library Journal/School Library Journal Previews, 1973 Includes three times a year Audiovisual guide previously issued as part of Library journal *Title Index, ERIC ED Accessions File* ERIC Processing and Reference Facility, 1966 Recording for the Blind & Dyslexic, ... Catalog of Books, 1996

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web so  $f = ma$  can be written as follows  $f = m \frac{dv}{dt}$  where  $f$  force since  $m \frac{dv}{dt}$  denotes a change in momentum  $f = m \frac{dv}{dt}$  if a force  $f$  is applied to a body

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