

## Stoichiometry Study Guide

Students should be able to...

- Explain [mole ratio](#)
- Convert moles, grams, atoms/molecules
- Calculate theoretical yields when given the limiting reactant (Convert: [moles A to moles B](#), [grams A to moles B](#), [moles A to grams B](#), [grams A to grams B](#))
- Explain [limiting reactant and excess](#).
- Calculate theoretical yields by finding the [limiting reactant](#)
- Calculate [percent yield](#).
- Calculate percent composition.
- Calculate theoretical yield, limiting reactant, excess, and percent yield for a double replacement reaction (laboratory)

### Practice Problems

Find the molar mass of each compound.

1.  $\text{BaI}_2$
2.  $\text{AgC}_2\text{H}_3\text{O}_2$

Convert.

3. 0.4 moles of Sn to atoms
4. 32 g of  $\text{K}_2\text{O}$  to molecules
5. 3.7 g of  $\text{H}_2$  to moles
6. 800 g of  $\text{CaCl}_2$  to moles
7. 0.75 moles of  $\text{H}_2\text{O}$  to grams
8. 9.0 g of  $\text{Cl}_2$  to atoms
9. 0.5 moles of  $\text{KNO}_3$  to grams

Find the percent composition for each element in the compound.

10.  $\text{Mn}_2\text{O}_7$
11.  $\text{AuNO}_3$

### Stoichiometry



If you started with 3.7 moles of magnesium oxide, how many moles of magnesium hydroxide would be produced?



If 2.8 grams of sodium were used for the reaction, and an excess of iodine was present, how many moles of sodium iodide would be produced?



# Unit 7chemistry Stoichiometry Study Guide Answers

**N Colangelo**



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