

- 1.0×10^{-14} is the equilibrium constant for the autoionization of water (K_w)
- Equilibrium occurs in all solutions (acid, base, neutral)
 - All aqueous solutions have a detectable concentration of both hydronium and hydroxide ions
 - The product of all of these ion concentrations is always K_w
 - $K_w = H_3O^{+} \cdot OH^{-} = 1.0 \times 10^{-14}$
- If concentration of hydroxide ion is known, to determine the pOH:
 - $pOH = 14 - pH$
- If pOH is known, to determine pH:
 - $pH = 14 - pOH$
- If the hydronium concentration is known, to determine the hydroxide ion concentration :
 - $(OH^{-}) = 1.0 \times 10^{-14} / (H_3O^{+})$
- Calculate pH from the hydronium concentration:
 - $pH = -\log (H_3O^{+})$
- If pH is known, to determine the hydronium concentration:
 - $(H_3O^{+}) = 10^{-pH}$

Buffer Solutions

- A weak acid or base mixed with its conjugate
- It resists a change in pH when H⁺ or OH⁻ is added
 - when H⁺ or OH⁻ is added to the buffer:
 - Decreases relative buffer component ratio
 - Works out the stoichiometry to find the new concentrations of buffer components
 - Use equilibrium constant to determine the new pH
- Buffer in acidic buffer
 - Weak acid + Conjugate Base
 - Weak Base + Conjugate Acid
 - Strong Weak Acid + Strong Base
 - Strong Weak Base + Strong Acid

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Minjie Lin

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