

- $1.0 \times 10^{-14}$  is the equilibrium constant for the autoionization of water ( $K_w$ )
- Equilibrium occurs in all solutions (acid, base, neutral)
  - All aqueous solutions have a detectable concentration of both hydronium and hydroxide ions
  - The product of all of these ion concentrations is always  $K_w$
  - $K_w = \text{H}_3\text{O}^{+*} \text{OH}^{-} = 1.0 \times 10^{-14}$
- If concentration of hydroxide ion is known, to determine the pOH:
  - $\text{pOH} = 14 - \text{pH}$
- If pOH is known, to determine pH:
  - $\text{pH} = 14 - \text{pOH}$
- If the hydronium concentration is known, to determine the hydroxide ion concentration :
  - $(\text{OH}^{-}) = 1.0 \times 10^{-14} / (\text{H}_3\text{O}^{+})$
- Calculate pH from the hydronium concentration:
  - $\text{pH} = \log (\text{H}_3\text{O}^{+})$
- If pH is known, to determine the hydronium concentration:
  - $(\text{H}_3\text{O}^{+}) = 10^{-\text{pH}}$



### Buffer Solutions

- A weak acid or base mixed with its conjugate
- It resists a change in pH when  $\text{H}^{+}$  or  $\text{OH}^{-}$  is added
- when  $\text{H}^{+}$  or  $\text{OH}^{-}$  is added to the buffer:
  - Determine which buffer component reacts
  - Work out the stoichiometry to find the new concentrations of buffer components
  - Use equilibrium constant to determine the new pH
- Ways to make buffers
  - Weak acid + Conjugate Base
  - Weak Base + Conjugate Acid
  - Excess Weak Acid + Strong Base
  - Excess Weak Base + Strong Acid

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**Minjie Lin**



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