

Chapter 1 • Introduction

P1.1 A gas at 200C may be *rarefied* if it contains less than 10¹² molecules per mm³. If Avogadro's number is 6.023E23 molecules per mole, what air pressure does this represent?

Solution: The mass of one molecule of air may be computed as

$$m = \frac{\text{Molecular weight}}{\text{Avogadro's number}} = \frac{28.97 \text{ mol}^{-1}}{6.023\text{E}23 \text{ molecules/g mol}} = 4.81\text{E}-23 \text{ g}$$

Then the density of air containing 10¹² molecules per mm³ is, in SI units,

$$= 4.81\text{E}-11 \frac{\text{g}}{\text{mm}^3} = 4.81\text{E}-5 \frac{\text{kg}}{\text{m}^3}$$

Finally, from the perfect gas law, Eq. (1.13), at 200C = 293 K, we obtain the pressure:

$$p = \rho RT = \left(4.81\text{E}-5 \frac{\text{kg}}{\text{m}^3} \right) \left(287 \frac{\text{m}^2}{\text{s}^2 \cdot \text{K}} \right) (293 \text{ K}) = 4.0 \text{ Pa} \quad \text{Ans.}$$

White Solution Manual

Michael Hoelscher



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