

Basic Chemistry Study Guide 2

1. Hydrogen bonds

1. Form between two polar molecules **covalent bonds**
2. A hydrogen's δ^+ region is attracted to a δ^- region in another molecule.
3. Form between water molecules (Where are the δ^+ and δ^- regions in a water molecule? **hydrogen bonds, at the top of the symbol**)
4. Each hydrogen bond is very weak (can be easily formed and broken).
5. There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life
6. One of these properties is the ability to be a solvent for other polar molecules.
7. Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

2. Solutions, solvents, and solutes

1. Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
2. Water is located inside and outside of our cells.
3. Body fluids are solutions
4. Define the terms **solution, solvent, and solute**. **Solvent: a substance capable of dissolving other polar molecules and ionic bonds, solutes: molecules dissolved in the solvent, solution: homogeneous mixture**

3. How substances dissolve in water

1. When a solute dissolves, it becomes surrounded by water molecules.
2. For a solute to dissolve, it must attract water molecules.
3. What kinds of substances attract water molecules? **Sugars, nucleic acids, and amino acids**
4. When a solute dissolves it forms hydrogen bonds with water molecules.
5. Define the terms **ion, cation, and anion**. **ion: atom with a charge, cation: positively charged ion, anions: negatively charged ion**
6. Define the terms **hydrophilic** and **hydrophobic** **Hydrophilic: attracted to water, hydrophobic: afraid of water**
7. What types of molecules are hydrophilic and what types are hydrophobic? **Nonpolar hydrophobic, polar hydrophilic**

Study Guide Chemistry Answers

Neil E. Schore



Study Guide Chemistry Answers:

Study Guide for Chemical Principles [by] Steven S. Zumdahl Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples

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Table of Contents Study Guide Chemistry Answers

1. Understanding the eBook Study Guide Chemistry Answers
 - The Rise of Digital Reading Study Guide Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Chemistry Answers
 - Personalized Recommendations
 - Study Guide Chemistry Answers User Reviews and Ratings

- Study Guide Chemistry Answers and Bestseller Lists
- 5. Accessing Study Guide Chemistry Answers Free and Paid eBooks
 - Study Guide Chemistry Answers Public Domain eBooks
 - Study Guide Chemistry Answers eBook Subscription Services
 - Study Guide Chemistry Answers Budget-Friendly Options
- 6. Navigating Study Guide Chemistry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Study Guide Chemistry Answers Compatibility with Devices
 - Study Guide Chemistry Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide Chemistry Answers
 - Highlighting and Note-Taking Study Guide Chemistry Answers
 - Interactive Elements Study Guide Chemistry Answers
- 8. Staying Engaged with Study Guide Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide Chemistry Answers
- 9. Balancing eBooks and Physical Books Study Guide Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide Chemistry Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide Chemistry Answers
 - Setting Reading Goals Study Guide Chemistry Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide Chemistry Answers
 - Fact-Checking eBook Content of Study Guide Chemistry Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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