

Limiting Reactant Practice

5 pages + 5-page Answer KEY + Periodic Table

chemistry

LIMITING REACTANTS

1 The reaction below produces ammonia for use in fertilizers.
 $N_2 + 3H_2 \rightarrow 2NH_3$
 If you start with 10 moles of N_2 and 30 moles of H_2 , determine the limiting reactant.

2 Calculate the grams of ammonia produced.

3 How many grams of the excess reactant is produced in a left over?
 To multiply the moles of the limiting reactant by the mole ratio of the excess reactant.

4 Determine moles of excess \rightarrow mass of excess

5 The reaction below produces carbon dioxide and water from glucose (sugar) and oxygen.
 $C_6H_{12}O_6 + 6O_2 \rightarrow 6CO_2 + 6H_2O$
 If you have 100 grams of glucose and 60 grams of oxygen, determine the limiting reactant.

6 How many grams of water (steam) are produced?



7 Magnesium chloride is used to create salts for de-icing roads. The reaction with very impure iron ore is produced.
 $Mg + 2FeCl_2 \rightarrow MgCl_2 + 2Fe$
 If you start with 10 moles of magnesium and 2 moles of iron(II) chloride, determine the limiting reactant.

8 Calculate the grams of magnesium chloride produced.

9 How many grams of the excess reactant is produced in a left over?

10 Magnesium chloride is used to create salts for de-icing roads. The reaction with very impure iron ore is produced.
 $Mg + 2FeCl_2 \rightarrow MgCl_2 + 2Fe$
 If you have with 100 grams of magnesium and 60 grams of iron(II) chloride, determine the limiting reactant.

11 How much carbon dioxide is produced?

12 The reaction below produces carbon dioxide and water from glucose (sugar) and oxygen.
 $C_6H_{12}O_6 + 6O_2 \rightarrow 6CO_2 + 6H_2O$
 If you have 100 grams of glucose and 60 grams of oxygen, determine the limiting reactant.

13 Calculate the grams of carbon dioxide produced.
 $100 \text{ g } C_6H_{12}O_6 \times \frac{1 \text{ mol } C_6H_{12}O_6}{180 \text{ g } C_6H_{12}O_6} \times \frac{6 \text{ mol } CO_2}{1 \text{ mol } C_6H_{12}O_6} \times \frac{44 \text{ g } CO_2}{1 \text{ mol } CO_2} = 155.56 \text{ g } CO_2$
 $60 \text{ g } O_2 \times \frac{1 \text{ mol } O_2}{32 \text{ g } O_2} \times \frac{6 \text{ mol } CO_2}{6 \text{ mol } O_2} \times \frac{44 \text{ g } CO_2}{1 \text{ mol } CO_2} = 66 \text{ g } CO_2$
 Since 66 g < 155.56 g, O_2 is the limiting reactant.

14 How many grams of the excess reactant is produced in a left over?
 To multiply the moles of the limiting reactant by the mole ratio of the excess reactant.

15 How many grams of carbon dioxide is produced?

16 Calculate the grams of carbon dioxide produced.
 $100 \text{ g } C_6H_{12}O_6 \times \frac{1 \text{ mol } C_6H_{12}O_6}{180 \text{ g } C_6H_{12}O_6} \times \frac{6 \text{ mol } CO_2}{1 \text{ mol } C_6H_{12}O_6} \times \frac{44 \text{ g } CO_2}{1 \text{ mol } CO_2} = 155.56 \text{ g } CO_2$
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 Since 66 g < 155.56 g, O_2 is the limiting reactant.

17 How many grams of the excess reactant is produced in a left over?
 To multiply the moles of the limiting reactant by the mole ratio of the excess reactant.

18 Calculate the grams of carbon dioxide produced.
 $100 \text{ g } C_6H_{12}O_6 \times \frac{1 \text{ mol } C_6H_{12}O_6}{180 \text{ g } C_6H_{12}O_6} \times \frac{6 \text{ mol } CO_2}{1 \text{ mol } C_6H_{12}O_6} \times \frac{44 \text{ g } CO_2}{1 \text{ mol } CO_2} = 155.56 \text{ g } CO_2$
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 Since 66 g < 155.56 g, O_2 is the limiting reactant.

19 How many grams of carbon dioxide is produced?

ANSWER KEY INCLUDED

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Tom Hsu PH.D.



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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. **Chemistry, Student Study Guide** James E. Brady, Fred

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Study Guide, Chemical Principles, Fifth Edition, Zumdahl Paul B. Kelter,Steven S. Zumdahl,2004-04

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Prentice Hall Chemistry, 2000 Understanding Chemistry Fred M. Dewey, 1994 **Essentials of Introductory Chemistry** Russo Steve Silver Michael, Steve Russo, 2001-12 Introductory Chemistry Third Edition helps readers master the quantitative skills and conceptual understanding they need to gain a deep understanding of chemistry Unlike other books on the market that emphasize rote memory of problem solving algorithms Introductory Chemistry takes a conceptual approach with the idea that focusing on the concepts behind chemical equations helps readers become more proficient problem solvers What Is Chemistry The Numerical Side of Chemistry The Evolution of Atomic Theory The Modern Model of the Atom 1 Chemical Bonding and Nomenclature The Shape of Molecules Chemical Reactions Stoichiometry and the Mole The Transfer of Electrons from One Atom to Another in a Chemical Reaction Intermolecular Forces and the Phases of Matter What If There Were No Intermolecular Forces The Ideal Gas Solutions When Reactants Turn into Products Chemical Equilibrium Electrolytes Acids and Bases For all readers interested in introductory chemistry **Foundations of College Chemistry, Alternate** Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis **Foundations of Physical Science** Tom Hsu PH.D., 2002 **Holt Chemistry** Ralph Thomas Myers, 2004

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