

balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds.

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium ($\text{H}-3$) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

61.5 yr

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} \boxed{12.5 \text{ mg}}$$

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20 \\ + & & \\ \hline & & 63.62 \rightarrow 64 \end{array}$$

Study Guide Answer Key Modern Chemistry

Lauren Gardner



Study Guide Answer Key Modern Chemistry:

Principles of Modern Chemistry Wade A. Freeman, David W. Oxtoby, 1996 PRINCIPLES OF MODERN CHEMISTRY has dominated the honors and high mainstream general chemistry courses and is considered the standard for the course. The fifth edition is a substantial revision that maintains the rigor of previous editions but reflects the exciting modern developments taking place in chemistry today. Authors David W. Oxtoby and H. P. Gillis provide a unique approach to learning chemical principles that emphasizes the total scientific process from observation to application, placing general chemistry into a complete perspective for serious-minded science and engineering students. Chemical principles are illustrated by the use of modern materials comparable to equipment found in the scientific industry. Students are therefore exposed to chemistry and its applications beyond the classroom. This text is perfect for those instructors who are looking for a more advanced general chemistry textbook. Student Study Guide/solutions Manual to Accompany Foundations of Inorganic, Organic, and Biological Chemistry Katherine J. Denniston, 1995. This student study guide has been written to accompany *Foundations of Inorganic Organic and Biological Chemistry* by Carey Denniston and Topping. It was designed to complement the text, not to be used in place of the text. Each chapter of the study guide contains the following sections: reorganized and expanded set of learning goals, concise chapter summary, in-chapter solved problems, list of key terms, review problems. Preface. Catalog of Copyright Entries. Third Series Library of Congress. Copyright Office, 1965. Includes Part 1 Number 2 Books and Pamphlets Including Serials and Contributions to Periodicals July-December. El-Hi Textbooks in Print, 1984. **Virginia Journal of Education**, 1958. El-Hi Textbooks and Serials in Print, 1995. R. R. Bowker Publishing, 1995-02. **U.S. Naval Training Bulletin**, 1948. The Cumulative Book Index, 1910. A world list of books in the English language. **The Education Gazette of the Province of the Cape of Good Hope** Cape of Good Hope (South Africa). Education Department, 1915.

The Virginia Journal of Education, 1958. **Barron's how to Prepare for College Entrance Examinations** Samuel C. Brownstein, Mitchel Weiner, 1974. A guide to preparing for college entrance examinations with emphasis on study programs for the verbal, mathematics, and standard written English parts of the SAT. Includes practice tests. **Books in Print Supplement**, 2002. The Science Teacher, 1967. Some issues are accompanied by a CD-ROM on a selected topic. Books and Pamphlets, Including Serials and Contributions to Periodicals Library of Congress. Copyright Office, 1973. **Books in Print**, 1981. **The Publishers' Trade List Annual**, 1887. **Naval Training Bulletin**, 1946. The Gregg Writer, 1912. **Today's Secretary**, 1912. Catalogue of Title-entries of Books and Other Articles Entered in the Office of the Librarian of Congress, at Washington, Under the Copyright Law ... Wherein the Copyright Has Been Completed by the Deposit of Two Copies in the Office Library of Congress. Copyright Office, 1978.

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