

Basic Chemistry Study Guide 2

1. Hydrogen bonds

1. Form between two polar molecules **covalent bonds**
2. A hydrogen's &+ region is attracted to a &- region in another molecule.
3. Form between water molecules (Where are the &+ and &- regions in a water molecule? **hydrogen bonds, at the top of the symbol**)
4. Each hydrogen bond is very weak (can be easily formed and broken).
5. There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life.
6. One of these properties is the ability to be a solvent for other polar molecules.
7. Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

2. Solutions, solvents, and solutes

1. Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
2. Water is located inside and outside of our cells.
3. Body fluids are solutions.
4. Define the terms **solution**, **solvent**, and **solute** **Solvent: a substance capable of dissolving other polar molecules and ionic bonds; solutes: molecules dissolved in the solvent, solution: homogeneous mixture**

3. How substances dissolve in water

1. When a solute dissolves, it becomes surrounded by water molecules.
2. For a solute to dissolve, it must attract water molecules.
3. What kinds of substances attract water molecules? **Sugars, nucleic acids, and amino acids**
4. When a solute dissolves it forms hydrogen bonds with water molecules.
5. Define the terms **ion**, **cation**, and **anion**. **ion: atom with a charge; cation: positively charged ion, anions: negatively charged ion**
6. Define the terms **hydrophilic** and **hydrophobic** **Hydrophilic: attracted to water; hydrophobic: afraid of water**
7. What types of molecules are hydrophilic and what types are hydrophobic? **Nonpolar: hydrophobic, polar: hydrophilic**

Study Guide Chemistry Bond Formations

McMurry, Susan McMurry

Study Guide Chemistry Bond Formations:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03

The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step by step examples. Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning. This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

Study Guide to Accompany Basics for Chemistry Martha Mackin, 1982-01-01 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook. Each chapter contains Overview, Topical Outline, Skills and Common Mistakes which are all keyed to the textbook for easy cross reference. The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms, a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook. The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples. The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter. After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter, energy measurement, problem solving, atom, periodic table and chemical bonding. These topics are followed by discussions on writing names and formulas of compounds, chemical formulas and the mole, chemical reactions, calculations based on equations, gases and the properties of a liquid. The remaining chapters examine the solutions, acids, bases, salts, oxidation-reduction reactions, electrochemistry, chemical kinetics and equilibrium and nuclear, organic and biological chemistry. This study guide will be of great value to chemistry teachers and students.

Basic Concepts of Chemistry, Study Guide Leo J. Malone, 1989-02-21 This Third Edition revised to provide smoother transitions between topics employs a concise yet informal approach to basic chemistry organized to help students employ basic math skills and problem solving strategies. Writing style is straightforward and presentation incorporates many concrete analogies to clarify new concepts. Includes

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