

# CHAPTER 8 STUDY GUIDE

## Covalent Bonding

### Section 8.1 The Covalent Bond

*In your textbook, read about the nature of covalent bonds.*

Use each of the terms below just once to complete the passage.

covalent bond

molecule

sigma bond

exothermic

pi bond

When sharing of electrons occurs, the attachment between atoms that results is called a(n) **(1)** \_\_\_\_\_. When such an attachment is formed, bond dissociation energy is released, and the process is **(2)** \_\_\_\_\_. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) **(3)** \_\_\_\_\_. If the electrons shared are centered between the two atoms, the attachment is called a(n) **(4)** \_\_\_\_\_. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) **(5)** \_\_\_\_\_.

*In your textbook, read about single and multiple bonds and bond strength.*

Circle the letter of the choice that best completes the statement or answers the question.

6. In what form do elements such as hydrogen, nitrogen, and oxygen normally occur?
  - a. as single atoms
  - b. as molecules containing two atoms
  - c. as molecules containing three atoms
  - d. as molecules containing four atoms
7. How many electrons are shared in a double covalent bond?
  - a. none
  - b. one
  - c. two
  - d. four
8. Bond length is the distance between
  - a. two molecules of the same substance.
  - b. the electrons in two attached atoms.
  - c. the nuclei of two attached atoms.
  - d. the orbitals of two attached atoms.
9. Which of the following relationships relating to bond length is generally correct?
  - a. the shorter the bond, the stronger the bond
  - b. the shorter the bond, the weaker the bond
  - c. the shorter the bond, the fewer the electrons in it
  - d. the shorter the bond, the lower the bond dissociation energy

# Study Guide Covalent Bonding

**Martha R. Taylor**



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## **Table of Contents Study Guide Covalent Bonding**

1. Understanding the eBook Study Guide Covalent Bonding
  - The Rise of Digital Reading Study Guide Covalent Bonding
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Covalent Bonding
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide Covalent Bonding
  - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Covalent Bonding
  - Personalized Recommendations
  - Study Guide Covalent Bonding User Reviews and Ratings
  - Study Guide Covalent Bonding and Bestseller Lists
5. Accessing Study Guide Covalent Bonding Free and Paid eBooks
  - Study Guide Covalent Bonding Public Domain eBooks
  - Study Guide Covalent Bonding eBook Subscription Services
  - Study Guide Covalent Bonding Budget-Friendly Options
6. Navigating Study Guide Covalent Bonding eBook Formats

- ePub, PDF, MOBI, and More
- Study Guide Covalent Bonding Compatibility with Devices
- Study Guide Covalent Bonding Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Study Guide Covalent Bonding
  - Highlighting and Note-Taking Study Guide Covalent Bonding
  - Interactive Elements Study Guide Covalent Bonding
- 8. Staying Engaged with Study Guide Covalent Bonding
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Study Guide Covalent Bonding
- 9. Balancing eBooks and Physical Books Study Guide Covalent Bonding
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Study Guide Covalent Bonding
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide Covalent Bonding
  - Setting Reading Goals Study Guide Covalent Bonding
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide Covalent Bonding
  - Fact-Checking eBook Content of Study Guide Covalent Bonding
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks



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