

# THE ULTIMATE GENERAL CHEMISTRY 1 STUDY GUIDE

## ❖ Matter and Measurement

- Matter: anything that has mass and occupies space
- Properties of matter: physical properties and chemical properties
- States of matter: solid, liquid, gas
- Measurements: SI Units, significant figures, dimensional analysis

## ❖ Atoms, Molecules, and Ions

- Atomic structure: protons, neutrons, and electrons
- Periodic table and trends: electronegativity, ionization energy, electron affinity
- Chemical bonding: ionic bonding, covalent bonding, metallic bonding
- Chemical formulas and equations: empirical and molecular formulas, balancing chemical equations

## ❖ Stoichiometry

- Limiting reactants and percent yield
- Mole concept and Avogadro's number
- Molarity and molality
- Gas laws: Boyle's law, Charles's law, Gay-Lussac's law, Avogadro's law, Ideal gas law

## ❖ Chemical Reactions and Equations

- Types of chemical reactions: synthesis, decomposition, single replacement, double replacement, combustion
- Redox reactions and balancing redox equations
- Reaction rates: factors affecting reaction rates, rate laws, and reaction mechanisms

## ❖ Gases

- Gas laws: Boyle's law, Charles's law, Gay-Lussac's law, Avogadro's law, Ideal gas law
- Kinetic Molecular Theory of gases: pressure, temperature, volume, and the number of moles of gas particles
- Natural gases: van der Waals equation

# Study Guide General Chemistry College

**Colorado Agricultural and Mechanical  
College. Department of  
Chemistry, Paul Reheard Frey**



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