

BALANCING EQUATIONS CHEAT SHEET

Steps for Balancing Chemical Equations

Step 1: Write the unbalanced chemical equation.

Step 2: Determine the number of atoms of each element in the unbalanced equation.

Step 3: Determine the smallest whole number ratio of reactants to products. If the ratio of the reactants to products is not a whole number, multiply the ratio by the least common multiple of the denominators of the ratios.

Step 4: Determine the balanced equation. Check to make sure the total number of atoms of each element in the balanced equation is the same as in the unbalanced equation. If the ratio of the reactants to products is not a whole number, multiply the ratio by the least common multiple of the denominators of the ratios.

Step 5: Check to make sure the balanced equation is balanced. If the balanced equation is not balanced, repeat the process from step 3.

Sample Equations



Step 1: Write the unbalanced chemical equation.

Reactants:

Products:

Step 2: Determine the number of atoms of each element in the unbalanced equation.

Reactants:

Products:

Step 3: Determine the number of atoms of each element in the balanced equation.

Reactants:

Products:

Step 4: Determine the balanced equation.

Step 5: Check to make sure the balanced equation is balanced.

Reactants:

Products:

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Pamphlet Master

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Chemical Reactions Science Learning Guide NewPath Learning, 2014-03-01 The Chemical Reactions Student Learning Guide includes self directed readings easy to follow illustrated explanations guiding questions inquiry based activities a lab investigation key vocabulary review and assessment review questions along with a post test. It covers the following standards aligned concepts Changes of Matter Chemical Reactions Formulas Balancing Equations Types of Chemical Reactions 1 Types of Chemical Reactions 2 Energy in Chemical Reactions Evidence of Chemical Reactions and Chemical Reaction Rates Catalysts Aligned to Next Generation Science Standards NGSS and other state standards *Study Guide with Selected Solutions* Karen Timberlake, 2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams. Also contains complete solutions to odd numbered problems *Study Guide to Accompany Chemical Principles, Properties, and Reactions* Kenneth L. Henold, 1984 *General Chemistry, Study Guide* James E. Brady, 1990-03-01 The Fifth Edition retains the pedagogical strengths that made the previous editions so popular and has been updated reorganized and streamlined. Changes include more accessible introductory chapters with greater stress on the logic of the periodic table

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