

Name _____ Date _____ Class _____

CHAPTER 15 **STUDY GUIDE FOR CONTENT MASTERY**

Solutions

Section 15.1 What are solutions?

In your textbook, read about the characteristics of solutions.

Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solution
insoluble	miscible	solvent	solvent

Air is air! (1) solution of oxygen gas dissolved in nitrogen gas. The oxygen in air is the (2) solute and nitrogen is the (3) solvent. Because oxygen gas dissolves in a solvent, oxygen gas is air's (4) soluble substance. A substance that does not dissolve is (5) insoluble. (6) Liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are (7) miscible. However, if one liquid is insoluble in another, the liquids are (8) immiscible.

Read about solvation in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



8. Hydration is solvation in which the solvent is water. What is solvation?

Solution is the process of surrounding solute particles with solvent particles to form a solution.

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CHAPTER 15 **STUDY GUIDE FOR CONTENT MASTERY**

Section 15.1 continued

10. As sodium chloride dissolves in water, what happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

11. Explain the attraction of the water molecules around the sodium ions and chloride ions.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

12. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.

13. List three ways that the rate of solvation may be increased.

stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent

In your textbook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write true or false.

true 14. The overall energy change that occurs when a solution forms is called the heat of solution.

false 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.

false 16. Solvation continues as long as the solvation rate is less than the crystallization rate.

true 17. In a saturated solution, solvation and crystallization are in equilibrium.

true 18. Additional solute can be dissolved in an unsaturated solution.

true 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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**Adams,Clark Edward Adams,Bernard J.
Nebel**

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notes definitions and chapter review questions i e short answer matching multiple choice and answer key Chapter topics

include 1 an overview of the human body 2 the problem of food choices 3 nutrient needs and nutrition surveys 4 energy

balance overweight and underweight 5 food composition and safety and 6 measures for improving one's diet The following nutrients are examined in detail carbohydrates sugar starch fiber lipids proteins vitamins minerals and water Nutrient requirements throughout the life cycle i.e. pregnancy infancy childhood adolescence adulthood are also reviewed Supplementary topics include the role of doctors in nutrition the appropriateness of dietary guidelines nutrition and acne and nutrition and cancer Appendices list reliable and unreliable sources of nutrition information and present a series of self study exercises and accompanying forms for evaluating individual dietary and exercise habits aje *Study Guide for Fundamentals of Nursing* Helen Harkreader, Mary Ann Hogan, Marshelle Thobaben, 2003-08 Practical and student friendly this helpful Study Guide is designed to assist students in mastering the content presented in the text Tied chapter by chapter to the text Includes Performance Checklists for all the procedures in the text Includes exercises on clinical judgment related to the case studies in the text chapter Includes chapter purpose learning objectives and a variety of learning exercises including matching true or false and fill in the blanks Answers to all activities and exercises are included in the back of the book Test Yourself sections present a series of NCLEX style multiple choice questions to provide practice in test taking and help with preparation for exams A multimedia focus with references to the main text to the text's EVOLVE website and to the Virtual Clinical Excursions that accompany this title Essentials of Pediatric Nursing Terri Kyle, 2008 Essentials of Pediatric Nursing is intended for Pediatric Nursing courses with an integrated pediatric curriculum It provides a unique concept based approach and nursing process focus that helps students go from concept to application by building on previously mastered knowledge from other courses Organized into four logical units Kyle Essentials of Pediatric Nursing covers a broad scope of topics with an emphasis on common issues and pediatric specific information In addition it has a variety of learning features to ensure student retention such as Healthy People 2010 boxes Threaded Case Studies and Comparison Charts highlighting common diseases Plus it includes a BONUS CD ROM and companion website that provide numerous resources for both students and instructors including video clips of each developmental stage and care of the hospitalized child **Telecourse Study Guide to The Business File** Raymond F. Attner, 1985

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