

Lesson 3: Limiting reactants

- Reactions stop when one of the reactants is used up.

Limiting reactant - is used completely in the reaction; determines the amount of product formed

Excess reactant - leftover once the reaction stops.

- To calculate which reactant is limiting, you will compare the amount of each reactant to a single product (the true amount of product formed came from the limiting)



If 200.0g of sulfur reacts with 100.0g of chlorine, what mass of disulfur dichloride is produced? which is the limiting reactant?

given = 200.0g S₈ ; 100.0g Cl₂

unknown = ?g S₂Cl₂ ; ? limiting reactant

$$200.0\text{g S}_8 \times \frac{1 \text{ mol S}_8}{256.528\text{g S}_8} \times \frac{4 \text{ mol S}_2\text{Cl}_2}{1 \text{ mol S}_8} \times \frac{135.038\text{g S}_2\text{Cl}_2}{1 \text{ mol S}_2\text{Cl}_2} = 421\text{g S}_2\text{Cl}_2$$

$$100.0\text{g Cl}_2 \times \frac{1 \text{ mol Cl}_2}{70.906\text{g Cl}_2} \times \frac{4 \text{ mol S}_2\text{Cl}_2}{1 \text{ mol Cl}_2} \times \frac{135.038\text{g S}_2\text{Cl}_2}{1 \text{ mol S}_2\text{Cl}_2} = \boxed{190\text{g S}_2\text{Cl}_2}$$

Chlorine is the limiting reactant



If 100.0g Na and 100.0g Fe₂O₃ react, what mass of Iron is produced? which is the limiting reactant?

given = 100.0g Na ; 100.0g Fe₂O₃

unknown = ?g Fe ; ? limiting reactant

$$100.0\text{g Na} \times \frac{1 \text{ mol Na}}{22.990\text{g Na}} \times \frac{2 \text{ mol Fe}}{6 \text{ mol Na}} \times \frac{55.845\text{g Fe}}{1 \text{ mol Fe}} = 81.0\text{g Fe}$$

$$100.0\text{g Fe}_2\text{O}_3 \times \frac{1 \text{ mol Fe}_2\text{O}_3}{159.687\text{g Fe}_2\text{O}_3} \times \frac{2 \text{ mol Fe}}{1 \text{ mol Fe}_2\text{O}_3} \times \frac{55.845\text{g Fe}}{1 \text{ mol Fe}} = \boxed{70.0\text{g Fe}}$$

L.R. = Fe₂O₃

Study Guide For Content Mastery Limiting Reactants

Ensheng Dong



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