

The Mole

Section 10.1 Measuring Matter

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Practice Problems

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1. Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel. Determine the number of Zn atoms in 2.50 mol of Zn.

$$2.50 \text{ mol-Zn} \times \frac{6.02 \times 10^{23} \text{ atoms-Zn}}{1 \text{ mol-Zn}} \\ = 1.51 \times 10^{24} \text{ atoms of Zn}$$

2. Calculate the number of molecules in 11.5 mol of water (H_2O).

$$11.5 \text{ mol-H}_2\text{O} \times \frac{6.02 \times 10^{23} \text{ molecules H}_2\text{O}}{1 \text{ mol-H}_2\text{O}} \\ = 6.92 \times 10^{24} \text{ molecules of H}_2\text{O}$$

3. Silver nitrate (AgNO_3) is used to make several different silver halides used in photographic films. How many formula units of AgNO_3 are there in 3.25 mol AgNO_3 ?

$$3.25 \text{ mol-AgNO}_3 \times \frac{6.02 \times 10^{23} \text{ formula units AgNO}_3}{1 \text{ mol-AgNO}_3} \\ = 1.96 \times 10^{24} \text{ formula units of AgNO}_3$$

4. Challenge Calculate the number of oxygen atoms in 5.0 mol of oxygen molecules. Oxygen is a diatomic molecule, O_2 .

$$5.00 \text{ mol-O}_2 \times \frac{6.02 \times 10^{23} \text{ molecules-O}_2}{1 \text{ mol-O}_2} \\ \times \frac{2 \text{ O atoms}}{\text{molecule-O}_2} = 6.02 \times 10^{24} \text{ atoms O}$$

5. How many moles contain each of the following?

a. 5.75×10^{24} atoms Al

$$5.75 \times 10^{24} \text{ atoms-Al} \times \frac{1 \text{ mol Al}}{6.02 \times 10^{23} \text{ atoms-Al}} \\ = 9.55 \text{ mol Al}$$

b. 2.50×10^{20} atoms Fe

$$2.50 \times 10^{20} \text{ atoms-Fe} \times \frac{1 \text{ mol Fe}}{6.02 \times 10^{23} \text{ atoms-Fe}} \\ = 4.15 \times 10^{-4} \text{ mol Fe}$$

6. Challenge Identify the representative particle for each formula and convert the given number of representative particles to moles.

a. $3.75 \times 10^{24} \text{ CO}_2$

The representative particle is a molecule.

$$3.75 \times 10^{24} \text{ molecules-CO}_2 \times \frac{1 \text{ mol CO}_2}{6.02 \times 10^{23} \text{ molecules-CO}_2} = 6.23 \text{ mol CO}_2$$

b. $3.58 \times 10^{23} \text{ ZnCl}_2$

The representative particle is a formula unit.

$$3.58 \times 10^{23} \text{ formula-units-ZnCl}_2 \times \frac{1 \text{ mol ZnCl}_2}{6.02 \times 10^{23} \text{ formula-units-ZnCl}_2} \\ = 0.595 \text{ mol ZnCl}_2$$

Section 10.1 Assessment

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7. Explain why chemists use the mole.

Chemists use the mole because it is a convenient way of knowing how many representative particles are in a sample.

8. State the mathematical relationship between Avogadro's number and 1 mol.

One mole contains 6.02×10^{23} representative particles.

9. List the conversion factors used to convert between particles and moles.

$$\frac{6.02 \times 10^{23} \text{ representative particles}}{1 \text{ mol}} \cdot \frac{1 \text{ mol}}{1 \text{ mol}} \cdot \frac{6.02 \times 10^{23} \text{ representative particles}}{6.02 \times 10^{23} \text{ representative particles}}$$

Study Guide The Mole Measuring Matter Answers

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Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon: buckyballs, graphite, and diamond are illustrated at the left as is the molecule methane CH₄ from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement.

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