

Chemistry: Atom Unit Study Guide

24. Fill in the blanks below for the following concepts:

25. Charge of the electron
26. Charge of the proton
27. Charge of the neutron
28. Charge of the alpha particle
29. Charge of the beta particle

25. Charge of the electron

26. Charge of the proton

27. Charge of the neutron

28. Charge of the alpha particle

29. Charge of the beta particle

Name: _____ Period: _____

Atom Unit Study Guide

Matching: Draw a line to match the words listed to the definitions.

- | | |
|------------------|---------------------------------------------------------------------------------|
| 1. Isotope | — Atoms of the same element with different numbers of neutrons |
| 2. Atomic Number | — The subscript particle has a positive charge |
| 3. Mass Number | — The subscript particle has a negative charge |
| 4. Proton | — The subscript particle has no charge (neutral) |
| 5. Neutron | — The number indicates the total number of protons and neutrons in an atom |
| 6. Electron | — This number indicates the number of protons in an atom and defines an element |

Circle the Blank

Word Bank: Atomic Number, Neutrons, Empty Mass, Neutrons, Protons, Protons Number (Don't use for words more than once)

- Where is most of an atom's mass found? _____
- An atom is made _____
- Which subatomic particle is the atomic number equal to? _____
- The atomic number is equal to _____ & _____ the combined subatomic the amount of Neutrons is _____
- Isotopes of the same atom have the same _____ number but different _____ numbers

Calculate

- Determine the number of protons, neutrons, and electrons of ^{14}C .
- How many neutrons does the element Silver (Ag) have?

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Daniel C. Pantaleo



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1. Understanding the eBook Study Guide The Atom Answers
 - The Rise of Digital Reading Study Guide The Atom Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide The Atom Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide The Atom Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide The Atom Answers

- Personalized Recommendations
- Study Guide The Atom Answers User Reviews and Ratings
- Study Guide The Atom Answers and Bestseller Lists
- 5. Accessing Study Guide The Atom Answers Free and Paid eBooks
 - Study Guide The Atom Answers Public Domain eBooks
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 - Study Guide The Atom Answers Budget-Friendly Options
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 - Highlighting and Note-Taking Study Guide The Atom Answers
 - Interactive Elements Study Guide The Atom Answers
- 8. Staying Engaged with Study Guide The Atom Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide The Atom Answers
- 9. Balancing eBooks and Physical Books Study Guide The Atom Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide The Atom Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide The Atom Answers
 - Setting Reading Goals Study Guide The Atom Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide The Atom Answers

- Fact-Checking eBook Content of Study Guide The Atom Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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