

# The Mole Study Guide

Topic	Definition	Example
Definition of a Mole	Avogadro's number of particles	6.02 x 10 <sup>23</sup>
Converting between moles and particles	$\frac{\text{moles}}{\text{Avogadro's number}} = \text{particles}$	1 mole of H <sub>2</sub> = 6.02 x 10 <sup>23</sup> H <sub>2</sub> molecules
Converting between moles and mass	$\frac{\text{moles}}{\text{molar mass}} = \text{mass}$	1 mole of H <sub>2</sub> = 2.016 g
Converting between mass and moles	$\frac{\text{mass}}{\text{molar mass}} = \text{moles}$	1 g of H <sub>2</sub> = 0.498 moles
Converting between moles and volume	$\frac{\text{moles}}{\text{molar volume}} = \text{volume}$	1 mole of H <sub>2</sub> = 22.4 L at STP
Converting between volume and moles	$\frac{\text{volume}}{\text{molar volume}} = \text{moles}$	1 L of H <sub>2</sub> = 0.0446 moles at STP
Converting between moles and energy	$\frac{\text{moles}}{\text{molar energy}} = \text{energy}$	1 mole of H <sub>2</sub> = 120 kJ
Converting between energy and moles	$\frac{\text{energy}}{\text{molar energy}} = \text{moles}$	1 kJ of H <sub>2</sub> = 0.00833 moles
Converting between moles and time	$\frac{\text{moles}}{\text{molar time}} = \text{time}$	1 mole of H <sub>2</sub> = 120 seconds
Converting between time and moles	$\frac{\text{time}}{\text{molar time}} = \text{moles}$	1 second of H <sub>2</sub> = 0.00833 moles
Converting between moles and space	$\frac{\text{moles}}{\text{molar space}} = \text{space}$	1 mole of H <sub>2</sub> = 120 meters
Converting between space and moles	$\frac{\text{space}}{\text{molar space}} = \text{moles}$	1 meter of H <sub>2</sub> = 0.00833 moles
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## Study Guide The Mole Answers

**Deitra Leonard Lowdermilk, Kitty  
Cashion, Kathryn Rhodes Alden, Ellen  
Olshansky, Shannon E. Perry**

## **Study Guide The Mole Answers:**

Study Guide with Answers to Selected Problems H. Stephen Stoker,Danny V. White,2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises *Chemistry, Student Study Guide* James E. Brady,Fred Senese,2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

**Study Guide with Selected Solutions** Karen Timberlake,2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems

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**Study Guide for General Chemistry and College Chemistry**, Eighth Editions by Holtzclaw and Robinson

**Norman E. Griswold, Henry Fuller Holtzclaw, 1988**

**Study Guide with Computer Exercises to Accompany Physics for Scientists and Engineers, Second Edition [and] Physics for Scientists and Engineers with Modern Physics, Second Edition** Raymond A. Serway, 1986

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