

Chapter 8 - Solution Stoichiometry - Multiple Choice Questions

1. Solubility is a term:
 a. that always means that more than 1g of solute will dissolve in 100 mL of water.
 b. that can be used for all solutes because they all dissolve to some extent.
 c. that generally means that at least 1g of solute will dissolve in 100 mL of water.
 d. that is used in the same way by all chemists.

2. Which statement is true?
 a. When an atom loses an electron the cation formed is smaller and when an atom gains an electron the anion formed is larger.
 b. When an atom gains or loses an electron the ion formed is smaller.
 c. When an atom gains or loses an electron the ion formed is larger.
 d. When an atom loses an electron the cation formed is larger and when an atom gains an electron the cation formed is smaller.

3. Two important factors that affect the solubility of an ionic compound are:
 a. ion charge and agitation.
 b. ion size and agitation.
 c. ion size and ion charge.
 d. ion size and the particle size of the compound.

4. Which statement is true?
 a. Compounds with small ions tend to be more soluble than compound with large ions.
 b. Compounds with large ions tend to be less soluble than compounds with small ions.
 c. Compounds with small ionic charges tend to be less soluble than compounds with large ionic charges.
 d. Compounds with large ions tend to be more soluble than compounds with small ions.

5. Which anion when combined with Mg^{2+} will probably give the most soluble ionic compound?
 a. F^- b. Cl^- c. Br^- d. I^-

6. Which cation when combined with oxide ion (O^{2-}) is most likely to give a soluble compound?
 a. Ca^{2+} b. Li^+ c. Al^{3+} d. Co^{2+}

7. According to the general solubility guidelines, which compound is not soluble?
 a. $CaCO_3$ b. $LiBr$ c. $CaCl_2$ d. $Mg(NO_3)_2$

8. According to the general solubility guidelines, which compound is soluble?
 a. $BaSO_4$ b. Na_2SO_4 c. Ag_2SO_4 d. $CuSO_4$

9. In a double displacement reaction:
 a. two ionic compounds react by exchanging ions.
 b. a precipitate may be formed.
 c. a gas may be produced.
 d. all of the above

10. Which combination of solutions will result in the formation of a precipitate?
 a. $Ba(NO_3)_2(aq) + Na_2CO_3(aq)$ c. $KNO_3(aq) + Na_2CO_3(aq)$
 b. $Ca(NO_3)_2(aq) + NaCl(aq)$ d. $Na_3NO_2(aq) + NaCl(aq)$

11. A metal carbonate reacts with an acid to produce a gas. Which gas is formed?
 a. $Na(g)$ b. $CO_2(g)$ c. $CO(g)$ d. $H_2O(g)$

12. Which of the following is one of the products of a reaction between an acid and a base?
 a. Water b. Hydrogen ion c. Hydroxide ion d. Hydrogen gas

13. Identify the spectator ions in the following total ionic equation:
 $Na^+(aq) + OH^-(aq) + Ag^+(aq) + NO_3^-(aq) \rightarrow Na^+(aq) + Ag(OH)_2 + NO_3^-(aq)$
 a. Na^+ and OH^- c. OH^- and Ag^+
 b. Ag^+ and $[NO_3^-]$ d. Na^+ and $[NO_3^-]$

14. The following equation is an example of what type of equation?
 $2Ag(s) + S^2(s) \rightarrow Ag_2S(s)$
 a. Molecular equation c. Net ionic equation
 b. Total ionic equation d. Skeletal equation

15. In writing the total ionic equation for the reaction between sodium phosphate and copper(II) chloride, which compound is not written as it's dissociated ion?
 a. Sodium chloride c. Copper(II) chloride
 b. Copper(II) phosphate d. Sodium phosphate

16. The net ionic equation for the reaction that occurs from mixing a solution of barium chloride with a solution of potassium sulfate is:
 a. $K^+(aq) + Cl^-(aq) \rightarrow KCl(s)$ c. $Ba^{2+}(aq) + SO_4^{2-}(aq) \rightarrow BaSO_4(s)$
 b. $BaCl_2(s) \rightarrow Ba^{2+}(aq) + 2Cl^-(aq)$ d. $BaSO_4(s) \rightarrow Ba^{2+}(aq) + SO_4^{2-}(aq)$

17. What is qualitative analysis?
 a. It determines how much of a compound, element or ion is in an unknown sample.
 b. It identifies compounds, elements or ions in an unknown solution.
 c. It measures concentrations of ions.
 d. None of the above.

18. What is the dissociation equation for an aqueous solution of magnesium chloride?
 a. $MgCl_2(aq) \rightarrow Mg^{2+}(aq) + Cl^-_2(aq)$ c. $MgCl_2(aq) \rightarrow Mg^{2+}(aq) + 2Cl^-(aq)$
 b. $MgCl_2(aq) \rightarrow Mg^{2+}(aq) + Cl^-(aq)$ d. $MgCl_2(aq) \rightarrow Mg(aq) + 2Cl^-(aq)$

19. What is the concentration of nitrate ions in a 0.25 mol/L solution of $Fe(NO_3)_3$?
 a. 0.45 mol/L
 b. 0.30 mol/L
 c. 0.15 mol/L
 d. There is not enough information to answer the question.

20. What is the concentration of a solution of aluminum sulfate ($Al_2(SO_4)_3$) if the concentration of sulfate ions is 0.50 mol/L?
 a. 1.2 mol/L b. 0.50 mol/L c. 0.40 mol/L d. 0.25 mol/L

21. Equal volumes of 0.20 mol/L $NaCl$ and 0.30 mol/L KCl are mixed together. What is the concentration of chloride ions in the mixture?
 a. 0.20 mol/L
 b. 0.25 mol/L
 c. 0.30 mol/L
 d. You need to know the volumes of the original solutions to answer the question.

22. What volume of a 0.500 mol/L solution of Na_2CO_3 contains enough carbonate ions to completely precipitate 0.0500 mol of Be^{2+} ions?
 a. 300 mL b. 90.0 mL c. 30.0 mL d. 10.0 mL

23. What volume of a 1.50 mol/L solution of $Be(OH)_2$ contains enough hydroxide ions to completely precipitate the iron(II) ions found in 60.0 mL of a 0.500 mol/L solution of $FeCl_2$?
 a. 180 mL b. 60.0 mL c. 20.0 mL d. 12.0 mL

24. How many moles of precipitate $BeSO_4$ are formed when 50.0 mL of a 0.500 mol/L solution of $BaCl_2$ is mixed with 25.0 mL of a 0.500 mol/L solution of Na_2SO_4 ?
 a. 0.0175 mol b. 0.0250 mol c. 0.0350 mol d. 0.0500 mol

25. How many moles of $Mg(OH)_2$ will be precipitated when 75.0 mL of 0.400 mol/L $MgCl_2$ is mixed with 75.0 mL of 0.400 mol/L $NaOH$?
 a. 0.0750 mol b. 0.0600 mol c. 0.0300 mol d. 0.0150 mol

26. In the above question, what mass of precipitate is formed?
 a. 0.875 g b. 1.75 g c. 3.50 g d. 4.37 g

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