

1.1 Find the rate of reaction defined as

$$r_{O_2} = \frac{\text{mol } O_2 \text{ used}}{\text{sec.} \cdot m^3 \text{ of tank}}$$

Evaluate terms

$$\bar{t} = \frac{V}{\bar{v}} \text{ or } V = \bar{t} \bar{v}$$

or (Volume of treatment tanks) $= \left(\frac{1}{3} \text{ day}\right) \left(32000 \frac{m^3}{\text{day}}\right) = 10667 m^3$

O_2 used

$$\left(200 \frac{\text{mg}}{\text{lit}}\right) \left(\frac{1 \text{ gm}}{1000 \text{ mg}}\right) \left(\frac{\text{mol}}{32 \text{ gm}}\right) \left(\frac{1000 \text{ lit}}{m^3}\right) \left(\frac{32000 m^3}{\text{day}}\right) = 2 \times 10^5 \frac{\text{mol } O_2}{\text{day}}$$

Thus the rate of reaction

$$\frac{2.0 \times 10^5 \text{ mol } O_2 / \text{day}}{10667 m^3} = 18.75 \text{ mol/m}^3 \cdot \text{day} \quad \left. \right\} \rightarrow$$

$$= 2.17 \times 10^4 \text{ mol/m}^3 \cdot \text{s}$$

1.2 Find $-r_{C_{20}H_{42}}$ and $-r_{C_{20}H_{42}}^m$ --- evaluate terms

$$V_{cat} = \frac{50000 \text{ kg}}{300 \text{ kg/m}^3} = 62.5 \text{ m}^3 \text{ of catalyst}$$

$$W_{cat} = 50000 \text{ kg}$$

$$m_w_{C_{20}H_{42}} = [20(12) + 42(1)] \frac{1}{1000} = 0.282 \text{ kg/mol}$$

$$F_{feed} = (6000 \text{ m}^3/\text{day}) (900 \text{ kg/m}^3) = 5400000 \text{ kg/day}$$

So $\frac{dN_{C_{20}H_{42}}}{dt} = \left(\frac{5400000 \text{ kg/day}}{0.282 \text{ kg/day}}\right) \left(\frac{\text{day}}{24(3600) \text{ s}}\right) (0.6) = 135 \text{ mol reacted/s}$

Thus the rate of disappearance of $C_{20}H_{42}$

$$-r' = \frac{1}{W_{cat}} \frac{dN}{dt} = \frac{1}{50000} (135) = 0.0027 \text{ mol/kg cat. s} \quad \left. \right\}$$

$$-r^m = \frac{1}{V_{cat}} \frac{dN}{dt} = \frac{1}{62.5} (135) = 2.16 \text{ mol/m}^3 \text{ cat. s} \quad \left. \right\}$$

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A new appendix has been added to summarize the developing technologies of bio treatment used to control contaminants
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