

1. Balance the following equations. Then write **NET** ionic equations for each. Assuming 1L of total solution, calculate the concentration of each ion left in solution after reaction.



2. Calculate the number of:

- (a) grams in 0.25 mol of magnesium hydroxide. (14.5 g)  
 (b) moles in 11.7 g of sodium chloride. (0.2 mol)  
 (c) moles in 4.9 g of hydrogen sulphate. (0.05 mol)

3. Calculate the molarity of solutions containing the following amounts of solute:

- (a) 31.5 g  $\text{HNO}_3$  per litre of solution. (0.5 M)  
 (b) 26.5 g  $\text{Na}_2\text{CO}_3$  per 200 mL of solution. (1.25 M)  
 (c) 66.57 g  $\text{AlCl}_3$  per 400 mL of solution. (1.25 M)

4. How many moles of solute are there in 100 mL of a 0.100 M solution? (0.01 mol)  
 5. How many grams of NaCl are there in 250 mL of a 2.50 M NaCl solution? (36.6 g)  
 6. What volume of a 0.75 M solution could be prepared from 500 g of  $\text{Na}_2\text{SO}_4$ ? (4.7 L)  
 7. One litre of 12 M HCl is diluted to 20 L. What is the molarity of the

# Solution Stoichiometry Practice Problems

**Silberberg**



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and relatable examples to demystify even the most challenging subjects in general chemistry

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**Study Guide, Chemical Principles, Fifth Edition, Zumdahl** Paul B. Kelter, Steven S. Zumdahl, 2004-04

**Introductory Chemistry** Charles H. Corwin, 2011 To help students learn chemical skills and concepts more effectively Introductory Chemistry Concepts and Critical Thinking Sixth Edition highlights the connection between key concepts and key problem solving skills through critical thinking Math and problem solving are covered early in the text Corwin builds your problem solving ability through innovative learning aids and technology formulated to meet your needs This revision retains

all the strengths of the previous editions while adding emphasis on conceptual understanding and critical thinking *The Software Encyclopedia 2000* Bowker Editorial Staff, 2000-05 **Introduction to Chemistry, Laboratory Manual** T. R. Dickson, 1994-12-23 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives Three themes are stressed throughout the text developing chemical thinking and a chemical vision learning problem solving methods and utilizing group work and discussion activities These themes involve and engage the students in their own learning processes they are challenged to be active The presentation of topics has been altered to include a new chapter which introduces the students to scientific thinking and shows that chemistry involves interesting and relevant topics The reorganization presents many core concepts in the first five chapters preparing students for later chapters In addition the author has added vignettes throughout the chapters referring to health technology the environment and society as well as to specific tools of direct use to students

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