

Basic Chemistry Study Guide 2

1. Hydrogen bonds

1. Form between two polar molecules **covalent bonds**
2. A hydrogen's &+ region is attracted to a &- region in another molecule.
3. Form between water molecules (Where are the &+ and &- regions in a water molecule? **hydrogen bonds, at the top of the symbol**)
4. Each hydrogen bond is very weak (can be easily formed and broken).
5. There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life.
6. One of these properties is the ability to be a solvent for other polar molecules.
7. Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

2. Solutions, solvents, and solutes

1. Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
2. Water is located inside and outside of our cells.
3. Body fluids are solutions.
4. Define the terms **solution**, **solvent**, and **solute** **Solvent: a substance capable of dissolving other polar molecules and ionic bonds; solutes: molecules dissolved in the solvent, solution: homogeneous mixture**

3. How substances dissolve in water

1. When a solute dissolves, it becomes surrounded by water molecules.
2. For a solute to dissolve, it must attract water molecules.
3. What kinds of substances attract water molecules? **Sugars, nucleic acids, and amino acids**
4. When a solute dissolves it forms hydrogen bonds with water molecules.
5. Define the terms **ion**, **cation**, and **anion**. **ion: atom with a charge; cation: positively charged ion, anions: negatively charged ion**
6. Define the terms **hydrophilic** and **hydrophobic** **Hydrophilic: attracted to water; hydrophobic: afraid of water**
7. What types of molecules are hydrophilic and what types are hydrophobic? **Nonpolar: hydrophobic, polar: hydrophilic**

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Table of Contents Test Study Guide Chemistry

1. Understanding the eBook Test Study Guide Chemistry
 - The Rise of Digital Reading Test Study Guide Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Test Study Guide Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Test Study Guide Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Test Study Guide Chemistry
 - Personalized Recommendations
 - Test Study Guide Chemistry User Reviews and Ratings
 - Test Study Guide Chemistry and Bestseller Lists
5. Accessing Test Study Guide Chemistry Free and Paid eBooks

- Test Study Guide Chemistry Public Domain eBooks
- Test Study Guide Chemistry eBook Subscription Services
- Test Study Guide Chemistry Budget-Friendly Options

6. Navigating Test Study Guide Chemistry eBook Formats

- ePUB, PDF, MOBI, and More
- Test Study Guide Chemistry Compatibility with Devices
- Test Study Guide Chemistry Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Test Study Guide Chemistry
- Highlighting and Note-Taking Test Study Guide Chemistry
- Interactive Elements Test Study Guide Chemistry

8. Staying Engaged with Test Study Guide Chemistry

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Test Study Guide Chemistry

9. Balancing eBooks and Physical Books Test Study Guide Chemistry

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Test Study Guide Chemistry

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Test Study Guide Chemistry

- Setting Reading Goals Test Study Guide Chemistry
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Test Study Guide Chemistry

- Fact-Checking eBook Content of Test Study Guide Chemistry
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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