

**Chapter 3**

1. An 18-g sample of element A combines completely with a 4-g sample of element B to form the compound AB. What is the mass of the compound formed?

$$\text{Mass}_{\text{reactants}} = \text{Mass}_{\text{products}}$$

$$\text{Mass}_A + \text{Mass}_B = \text{Mass}_{AB}$$

$$\text{Mass}_{AB} = 18 \text{ g} + 4 \text{ g} = 22 \text{ g}$$

2. A substance breaks down into three component elements when it is heated. The mass of each component element is listed in the table below. What was the mass of the substance before it was heated?

Component	Mass (g)
A	39.10
B	54.94
C	64.00

$$\text{Mass}_{\text{reactants}} = \text{Mass}_{\text{products}} = 39.10 + 54.94 + 64.00 = 158.04 \text{ g}$$

3. Silver iodide powder has been used as an antiseptic and as an agent to seed clouds for rain. Silver iodide is 45.9% silver by mass. If you separate a 50.0-g sample of silver iodide into its elements, silver and iodine, how much silver would you have?

From the conservation of mass, the mass of silver recovered is equal to the mass of silver in the initial silver iodide sample. The amount of silver recovered would be  $50.0 \text{ g} \times 45.9\% = 50.0 \times 0.459 = 22.95 \text{ g} = 23 \text{ g}$ .

4. If 5 g of element A combines with 16 g of element B to form compound AB, how many grams of B are needed to form compound AB<sub>2</sub>? How many grams of B are needed to form AB<sub>3</sub>?

Compound AB<sub>2</sub> contains twice as much element B as does compound AB. Therefore,  $2 \times 6 \text{ g} = 32 \text{ g}$ .

There is 32 g of B in AB<sub>2</sub>. Compound AB<sub>3</sub> contains three times as much element B as does compound AB. Therefore,  $3 \times 16 \text{ g} = 48 \text{ g}$ . There is 48 g of B in AB<sub>3</sub>.

5. During a chemical reaction, 2.445 g of carbon reacts with 3.257 g of oxygen to form carbon monoxide gas. How many grams of carbon monoxide are formed in this reaction?

$$\text{Mass}_{\text{reactants}} = \text{Mass}_{\text{products}}$$

$$\text{Mass}_{\text{carbon}} + \text{Mass}_{\text{oxygen}} = \text{Mass}_{\text{carbon monoxide}}$$

$$2.445 \text{ g} + 3.257 \text{ g} = 5.702 \text{ g}$$

6. Ibuprofen has the chemical formula C<sub>13</sub>H<sub>18</sub>O<sub>2</sub>. It is 75.69% carbon, 8.80% hydrogen, and 15.51% oxygen. How many mg of carbon does a 200-mg tablet of ibuprofen contain?

Mass percentage of an element (%):

$$\frac{\text{Mass of element}}{\text{Mass of compound}} \times 100\%$$

$$\text{Mass percentage}_{\text{carbon}} = \frac{\text{Mass}_{\text{carbon}}}{\text{Mass}_{\text{compound}}} \times 100\%$$

$$75.69\% \text{ carbon} = \frac{\text{Mass}_{\text{carbon}}}{200 \text{ mg}} \times 100\%$$

$$75.69\% \text{ carbon} \times \frac{200 \text{ mg}}{100\%} = \text{Mass}_{\text{carbon}}$$

$$\text{Mass}_{\text{carbon}} \text{ in the tablet} = 151.38 \text{ mg}$$

7. During a chemical reaction, 4.032 g of hydrogen combined with oxygen to form 36.032 g of water. How many grams of oxygen reacted?

$$\text{Mass}_{\text{reactants}} = \text{Mass}_{\text{products}}$$

$$\text{Mass}_{\text{hydrogen}} + \text{Mass}_{\text{oxygen}} = \text{Mass}_{\text{water}}$$

$$4.032 \text{ g} + \text{Mass}_{\text{oxygen}} = 36.032 \text{ g}$$

$$\text{Mass}_{\text{oxygen}} = 36.032 \text{ g} - 4.032 \text{ g} = 32 \text{ g}$$

# Supplemental Practice Problems Solutions Manual

## Chemistry

**Morris Hein**

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## **Table of Contents Supplemental Practice Problems Solutions Manual Chemistry**

1. Understanding the eBook Supplemental Practice Problems Solutions Manual Chemistry
  - The Rise of Digital Reading Supplemental Practice Problems Solutions Manual Chemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying Supplemental Practice Problems Solutions Manual Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook Supplemental Practice Problems Solutions Manual Chemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Supplemental Practice Problems Solutions Manual Chemistry
  - Personalized Recommendations
  - Supplemental Practice Problems Solutions Manual Chemistry User Reviews and Ratings

- Supplemental Practice Problems Solutions Manual Chemistry and Bestseller Lists
- 5. Accessing Supplemental Practice Problems Solutions Manual Chemistry Free and Paid eBooks
  - Supplemental Practice Problems Solutions Manual Chemistry Public Domain eBooks
  - Supplemental Practice Problems Solutions Manual Chemistry eBook Subscription Services
  - Supplemental Practice Problems Solutions Manual Chemistry Budget-Friendly Options
- 6. Navigating Supplemental Practice Problems Solutions Manual Chemistry eBook Formats
  - ePub, PDF, MOBI, and More
  - Supplemental Practice Problems Solutions Manual Chemistry Compatibility with Devices
  - Supplemental Practice Problems Solutions Manual Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Supplemental Practice Problems Solutions Manual Chemistry
  - Highlighting and Note-Taking Supplemental Practice Problems Solutions Manual Chemistry
  - Interactive Elements Supplemental Practice Problems Solutions Manual Chemistry
- 8. Staying Engaged with Supplemental Practice Problems Solutions Manual Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Supplemental Practice Problems Solutions Manual Chemistry
- 9. Balancing eBooks and Physical Books Supplemental Practice Problems Solutions Manual Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Supplemental Practice Problems Solutions Manual Chemistry
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Supplemental Practice Problems Solutions Manual Chemistry
  - Setting Reading Goals Supplemental Practice Problems Solutions Manual Chemistry
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Supplemental Practice Problems Solutions Manual Chemistry
  - Fact-Checking eBook Content of Supplemental Practice Problems Solutions Manual Chemistry
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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