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**Lingjun Ying**

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we start with 8.3 mole H<sub>2</sub>O and multiply it with 2 mole H<sub>2</sub>O over 2 mole H<sub>2</sub>O from the equation 8.3 mole H<sub>2</sub>O × 2 mole H<sub>2</sub>O / 2 mole H<sub>2</sub>O × 18 g H<sub>2</sub>O is your theoretical yield so percent yield 137.1 g H<sub>2</sub>O actual yield 149.4 g H<sub>2</sub>O × 100 / 91.77 problems 6.8 are similar to problem 5

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web a write a balanced equation for the reaction Al + O<sub>2</sub> → Al<sub>2</sub>O<sub>3</sub> 4 Al + 3 O<sub>2</sub> → 2 Al<sub>2</sub>O<sub>3</sub> b determine the theoretical yield of Al<sub>2</sub>O<sub>3</sub> 160.0 g Al × 1 mol Al × 2 mol Al<sub>2</sub>O<sub>3</sub> × 101.964 g Al<sub>2</sub>O<sub>3</sub> 302.3 g Al<sub>2</sub>O<sub>3</sub> 26.982 g Al × 4 mol Al × 1 mol Al<sub>2</sub>O<sub>3</sub> c determine the percent yield yield actual × 100 / 260.0 g × 100 / 86.01 theoretical 302.3 g

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web 7 Na + 2 C + O<sub>2</sub> → 4 Na<sub>2</sub>CO<sub>3</sub> only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H<sub>2</sub> are left over 13

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