

# Stoichiometry Study Guide

Answers the following questions. Then use the following equations to answer questions 3-8.

1.  $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$   
2.  $\text{Ca} + \text{O}_2 \rightarrow \text{CaO}$   
3.  $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$   
4.  $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$   
5.  $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$   
6.  $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$   
7.  $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$   
8.  $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$

Name: \_\_\_\_\_ Date: \_\_\_\_\_

### Stoichiometry Study Guide

Part A: Stoichiometry Questions

1.  $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$

Answers the following questions. Then use the balanced equation to answer the questions 3-8.

$$\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$$

1. Predict the balanced equation for the following reaction:

- $\text{MgO} + \text{O}_2$
- $\text{MgO} + \text{H}_2\text{O}$
- $\text{MgO} + \text{H}_2$

2. If the above chemical reaction is done as an experiment using 10 moles of each reactant, calculate how much product will be produced?

- Using 10 moles of  $\text{MgO}$ , how many moles of  $\text{MgO}$  will be produced?
- Using 10 moles of  $\text{MgO}$ , how many moles of  $\text{MgO}$  will be consumed?
- Using 10 moles of  $\text{MgO}$ , how many moles of  $\text{MgO}$  will be consumed?
- What is the limiting reactant in this reaction?

3. If the above chemical reaction is done as an experiment using 100 moles of  $\text{O}_2$ , calculate the percent yield for the actual yield of  $\text{MgO}$  to 100 moles.

- What is the theoretical yield in grams of  $\text{MgO}$  given 100 moles of  $\text{O}_2$ ?
- What is the percent yield given that the actual yield of  $\text{MgO}$  is 80 moles?

# Stoichiometry Unit Study Guide

**Martha B. Barrett**

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