

Stoichiometry Study Guide

Balance the following equations. Then use the balanced equation to answer questions #1-5.

1) 10 moles of H_2O are

a) 180 grams

b) 360 grams

c) 540 grams

2) 10 moles of H_2O are

a) 180 grams

b) 360 grams

c) 540 grams

3) 10 moles of H_2O are

a) 180 grams

b) 360 grams

c) 540 grams

Balance the following equation.

4) 10 moles of H_2O are

a) 180 grams

b) 360 grams

c) 540 grams

5) 10 moles of H_2O are

a) 180 grams

b) 360 grams

c) 540 grams

6) 10 moles of H_2O are

a) 180 grams

b) 360 grams

c) 540 grams

Name _____ Period _____ Date _____

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Fill in the blanks below.

1 mole = _____ grams.

Balance the following equation. Then use the balanced equation to answer the questions #1-5.



1) Find the formula for the following compounds:

a) NaCl to O_2

b) NaClO_2 to NaClO_2

c) O_2 to NaClO_2

2) If the above chemical reaction is done as an experiment using 4.4 moles of each reactant, calculate how much NaClO_2 will be produced?

a) Given 4.4 moles of NaCl , how many moles of NaClO_2 will be produced?

b) Given 4.4 moles of O_2 , how many moles of NaClO_2 will be produced?

c) Define limiting reactant & excess reactant.

d) Which reactant is the limiting reactant? Which reactant is the excess reactant?

3) If the above chemical reaction is done as an experiment using 10 moles of O_2 , calculate the percent yield given that the actual yield of NaClO_2 is 8.8 moles.

a) What is the theoretical yield in grams of NaClO_2 given 10 moles of the O_2 ?

b) Write the formula for Percent Yield.

c) What is the percent yield given that the actual yield of NaClO_2 is 8.8 moles?

Stoichiometry Unit Study Guide

Martha B. Barrett



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