

Stoichiometry Study Guide KEY

Chemistry

RHS – Mr. Moss

1. Define the following:
 - a. *Stoichiometry-the study of the quantitative relationships between the amounts of reactants used and the products formed by a chemical reaction.*
 - b. *Mole –The SI unit used to measure the amount of a substance that contains 6.02×10^{23} atoms of that substance.*
 - c. *Mole Ratio-The ratio between any two substances in a balanced chemical equation.*
 - d. *Molar Mass-The mass in grams of one mole of any substance.*
2. What relationships can be determined from a balanced chemical equation? *The relationships between particles, moles, and mass for all reactants and products.*
3. Explain why mole ratios are central to stoichiometric calculations. *Mole ratios allow for the conversion from moles of one substance in a balanced chemical equation to moles of another substance in the same equation.*
4. What is the first step in all stoichiometric calculation? (Be careful.) *Write a balanced chemical equation.*
5. How is molar mass used in some stoichiometric calculations? *Molar mass is a conversion factor for converting moles of a given substance to mass or mass of a given substance to moles.*
6. Find the molar mass of the following:
 - a. H_2SO_4 (sulfuric acid) *98.09 g/mol*
 - b. ZnO (zinc oxide) *81.39 g/mol*
 - c. NH_4NO_3 (ammonium nitrate) *80.05 g/mol*
7. List the mole ratios of the following *balanced* chemical equation.
 - a. $2\text{Al} + 3\text{Br}_2 \rightarrow 2\text{AlBr}_3$ $\frac{2 \text{ mol Al}}{3 \text{ mol Br}_2} ; \frac{3 \text{ mol Br}_2}{2 \text{ mol AlBr}_3} ; \frac{2 \text{ mol Al}}{2 \text{ mol AlBr}_3}$ *(Each one of these can be inverted to form another mole ratio. This would give a total of 6 ratios.)*

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Leo J. Malone, Theodore O. Dolter



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remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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