

# Chemistry: Atom Unit Study Guide

24. Fill in the boxes below for the following concepts:

1. Isotope
2. Atomic Number
3. Mass Number
4. Proton
5. Neutron
6. Electron

25. Draw a Bohr model of the element

26. Complete the following table:

1. Isotope
2. Atomic Number
3. Mass Number
4. Proton
5. Neutron
6. Electron

27. Draw a Bohr model of the element

1. Isotope
2. Atomic Number
3. Mass Number
4. Proton
5. Neutron
6. Electron

28. Using the periodic table, determine the number of protons, neutrons, and electrons of  $^{12}\text{C}$ .

Name: \_\_\_\_\_ Period: \_\_\_\_\_

## Atom Unit Study Guide

**Matching:** Draw a line to match the words listed to the definitions.

- |                  |  |
|------------------|--|
| 1. Isotope       | — Atoms of the same element with different numbers of neutrons.                  |
| 2. Atomic Number | — The subscript particle has a positive charge.                                  |
| 3. Mass Number   | — The subscript particle has a negative charge.                                  |
| 4. Proton        | — The subscript particle has no charge (neutral).                                |
| 5. Neutron       | — The number indicates the total number of protons and neutrons in an atom.      |
| 6. Electron      | — This number indicates the number of protons in an atom and defines an element. |

### Fill in the Blank:

**Word Bank:** Isotope, Neutrons, Empty Space, Neutrons, Protons, Protons, Neutrons  
(Words can be used more than once)

- Where is most of an atom's mass found? \_\_\_\_\_
- An atom is mostly \_\_\_\_\_.
- Which subatomic particle is the atomic number equal to? \_\_\_\_\_
- The atomic number is equal to \_\_\_\_\_, \_\_\_\_\_, the number indicates the amount of \_\_\_\_\_.
- Isotopes of the same atom have the same \_\_\_\_\_ number but different \_\_\_\_\_.

### Calculate:

- Determine the number of protons, neutrons, and electrons of  $^{12}\text{C}$ .
- How many neutrons does the element Silver (Ag) have?

# Study Guide The Atom Answers

**Robinson**



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