

The Mole Study Guide

Mass Conversions: Round your answers to 2 places after the decimal! Make sure to show your steps!

1) There are 3.75 moles of CO_2 .

a) mass

b) mass

c) mass

2) There are 2.5 moles of H_2O .

a) mass

b) mass

c) mass

3) There are 4.5 moles of H_2O .

a) mass

b) mass

4) There are 3.5 moles of H_2O .

a) mass

b) mass

5) There are 2.5 moles of H_2O .

a) mass

b) mass

Name _____ Period _____ Date _____

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Fill in the blank boxes for the image below.



Write the Empirical Formula for each compound.
(Use subscripts to show the relative ratio of atoms.)

Molecular	Empirical
$\text{C}_6\text{H}_{12}\text{O}_6$	
$\text{C}_4\text{H}_8\text{O}_4$	
$\text{C}_6\text{H}_{12}\text{O}_6$	
$\text{C}_6\text{H}_{12}\text{O}_6$	
$\text{C}_6\text{H}_{12}\text{O}_6$	

Calculate the Molar Mass for the following compounds.

1) H_2O

2) H_2O

3) H_2O

4) H_2O

Calculate the Percent Composition for each element in the following compound.

1) H_2O

2) H_2O

3) H_2O

4) H_2O

5) H_2O

6) H_2O

Determine the Empirical and Molecular Formula for the compound given the information below.

1) A compound has the following percent composition: 58.80% Hg, 21.80% Cl, 1.40% H. The molar mass for Aqueous is 100.0 g/mol.

a) If you have 100 g of Aqueous, how many grams of Hg, Cl, and H do you have?

b) Determine the Empirical Formula for Aqueous.

c) Determine the Molecular Formula for Aqueous.

Study Guide The Mole

Robert K. Wismer



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