

Molecular Polarity Table

A molecule is **non-polar** if all surrounding atoms are same.

A molecule is **polar**, if surrounding atoms are different.

A molecule is **polar**, if central atom has one or more lone pair of electrons.

BP: Bond pair

LP: Lone pair

A: Central atom

X: Surrounding atoms



Number of electron groups	Type of electron pairs	VSEPR notation	Name of Molecular shape	Example	Polarity
2	2 BP	AX_2	$X - A - X$ linear	BeF_2	Non-polar
3	3 BP	AX_3	 trigonal planar	BF_3	Non-polar
3	2 BP, 1 LP	AX_2E	 angular	$SnCl_3$	Polar
4	4 BP	AX_4	 tetrahedral	CF_4	Non-polar
4	3 BP, 1 LP	AX_3E	 trigonal bipyramidal	PCl_3	Polar
4	2BP, 2LP	AX_2E_2	 angular	H_2S	Polar

Study Guide Molecular Shape Answer Key

**David L. Nelson, Albert L.
Lehninger, Michael M. Cox**

Study Guide Molecular Shape Answer Key:

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