

**INSTRUCTIONS** (Read carefully): Carefully separate all sheets. This examination has 2 pages of problems, 1 double-sided answer sheet, a periodic table, and 1 blank sheet for scratch work. Read each question carefully. For 2 bonus points, correctly complete as requested on *both sides* of the answer sheet: print your name, sign your name, give the portion of your PSU e-mail address that appears before the @ (this is the same as your BlackBoard login user name), give your row letter and seat number, and give your Discussion Section number. Put your answers in the spaces provided on the answer sheet. *All sheets will be collected; only answers on the answer sheets will be graded.* There is a total of 100 points plus 2 bonus points. Before the next lecture, go over problems about which you were not confident on this examination.

**REGRADE POLICY:** If you have a question about the number of points that you earned on your graded exam answer sheet, *do not make changes or additions to your returned, graded answer sheet.* Write on a *separate sheet of paper* what items you believe that you were not given sufficient credit and the reasons why you think so. Attach it to your original answer sheet and give it to Dr. Hilinski after a lecture or during an office hour, preferably within 2 weeks after it is returned. Be sure to make and keep a photocopy for your records.

- [9 points] For the compound given to the right, provide a name.
- [12 points] For each of the following, consider the compound given in **problem 1**.
  - How many methylene hydrogens are there?
  - How many tertiary Cs are there?
  - Give the molecular formula of this compound.
  - For the *number* of Cs in this compound, what is the maximum number of Hs? Do not consider only the specific structural formula given in **problem 1**; do consider the general case for the number of Cs that you gave in part 2c.
  - How many tertiary Hs are there?
  - How many primary alkyl groups exist as substituents?
- [7 points] Provide a structural formula for 5-ethyl-2,3-dimethyl-8-(1-methylcyclopentyl)nonane. *For this problem, you must give a complete structural formula showing all atoms and all bonding valence electrons.*
- [22 points] Consider the following.



(a) One of the 5 resonance structures (Lewis structures) that contribute to the resonance hybrid of the conjugate base of phenol is represented above as one of the products in the given chemical equation. Draw the additional 4 resonance structures (Lewis structures) that contribute to the resonance hybrid of the conjugate base of phenol.

*Note that if you incorrectly duplicate resonance structures, you will lose credit for that resonance structure. If you give more than 4 resonance structures, points will be deducted.*

For each of the four Lewis structures (resonance structures) that you give, you must show all atoms, all bonding valence electrons, all unshared valence electrons, and each nonzero formal charge.

(b) The equilibrium shown above lies to the right. This, of course, means that phenol is more acidic than cyclohexanol. Using structural considerations, provide a clear, concise explanation (but in as much detail as possible) why this equilibrium lies to the right. *NOTE that a correct answer to this part is *not* that "the pK<sub>a</sub> of phenol is less than the pK<sub>a</sub> of cyclohexanol" because this is merely a restatement of the fact that phenol is more acidic than cyclohexanol.*

- [10 points] Draw structural formulas for 5 (*and only 5*) constitutional isomers that *have one (and only one) ring* and the molecular formula C<sub>4</sub>H<sub>6</sub>FO. Yes, each of the constitutional isomers that you give must be *have one (and only one) ring*. Your structural formulas must have no nonzero formal charges and the most typical arrangements of bonding and nonbonding valence electrons. *Note that if you give the same isomer more than once, points will be deducted for that isomer. If you give more than 5 constitutional isomers, points will be deducted. For this problem, you must give complete structural formulas showing all atoms and bonding valence electrons. Do not give condensed structural formulas or bond-line formulas.*
- [4 points] A condensed structural formula for allene is shown to the right. Using lines, wedges, and dashed wedges for bonding electrons, give a 3-D stereoformula for allene that is a representation of the true shape of this molecule.  
For your structure, label clearly each C with its hybridization (*sp*<sup>3</sup>, *sp*<sup>2</sup>, or *sp*).



# Study Guide Organic Chemistry Answers

**Ralph J. Fessenden,Joan S. Fessenden**

## **Study Guide Organic Chemistry Answers:**

**Study Guide and Solutions Manual for Essential Organic Chemistry** Paula Yurkanis Bruice,2015-03-04 This Study Guide Solution Manual contains learning objectives chapter summaries and outlines as well as examples self tests and concept questions as well as complete step by step solutions to selected problems

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Study Guide, Answer Book to Accompany Organic Chemistry Alan S. Wingrove,Robert L. Caret,1981 Study Guide and Solutions Manual for Organic Chemistry Digital Update K. Peter C. Vollhardt,Neil E. Schore,2024-03-25 Written by Organic Chemistry coauthor Neil Schore this invaluable manual includes chapter introductions that highlight new materials chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer *A Self-study Guide to the Principles of Organic Chemistry* Jiben Roy,2013 A Self Study Guide to the Principles of Organic Chemistry Key Concepts Reaction Mechanisms and Practice Questions for the Beginner will help students new to organic chemistry grasp the key concepts of the subject quickly and easily as well as build a strong foundation for future study Starting with the definition of atom the author explains molecules electronic configuration bonding hydrocarbons polar reaction mechanisms stereochemistry reaction varieties organic spectroscopy aromaticity and aromatic reactions biomolecules organic polymers and a synthetic approach to organic compounds The over one hundred diagrams and charts contained in this volume will help students visualize the structures and bonds as they read the text and make the logic of organic chemistry clear and easily understood Each chapter ends with a list of frequently asked questions and answers followed by additional practice problems Answers are included in the Appendix Organic Chemistry Study Guide with Solutions Manual Neil E. Schore,2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer

**Organic Chemistry, Study Guide and Solutions Manual** T. W. Graham Solomons,Craig B. Fryhle,2000 On the cover of this book is a Pacific yew tree found in the ancient forests of the Pacific Northwest The bark of the Pacific yew tree produces Taxol found to be a highly effective drug against ovarian and breast cancer Taxol blocks mitosis during eukaryotic cell division The supply of Taxol from the Pacific yew tree is vanishingly small however A single 100 year old tree provides only about one dose of the drug roughly 300 mg For this reason as well as the spectacular molecular architecture of Taxol synthetic organic chemists fiercely undertook efforts to synthesize it Five total syntheses of Taxol have thus far been

reported Now a combination of isolation of a related metabolite from European yew needles and synthesis of Taxol from that intermediate supply the clinical demand This case clearly demonstrates the importance of synthesis and the use of organic chemistry It's just one of the many examples used in the text that will spark the interest of students and get them involved in the study of organic chemistry

**Organic Chemistry Study Guide and Solutions** Marc Loudon, Jim Parise, 2015-07-01  
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Manual and Study Guide to Accompany Introduction to Organic Chemistry, 4th Ed Paul A. Bartlett, Judith G. Koch, 1992

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**Study Guide and Solutions Manual for Organic Chemistry** Susan McMurry, John McMurry, 1988 Study Guide with Solutions for Organic Chemistry, Third Edition Ralph J. Fessenden, Joan S. Fessenden, 1986 Provides a set of additional drill problems chapter by chapter discussions and supplemental instructional material to help students master organic chemistry problem solving techniques

**Study Guide and Additional Drill Problems for Organic Chemistry** Ralph J. Fessenden, 1990

**Study Guide & Solutions Manual** Paula Yurkanis Bruice, 2011 Study Guide and Solutions Manual for Fundamentals of Organic Chemistry Susan McMurry, 1994 Following a brief review of structure and bonding organic molecules and functional groups are presented as early as possible. The text is organized primarily by functional group beginning with simple alkanes and moving toward more complex compounds. Emphasis is placed on the fundamental mechanistic similarities of organic reactions. McMurry's thorough revision continues to present the solid content necessary for this course without sacrifice of important subjects and pedagogical tools. Text and reaction summaries full problem sets and outstanding artwork are just some of the features in the Third Edition usually found in a full year book. McMurry's clear well written explanations remain a highlight of the book

Study Guide and Solutions Manual to Accompany The Essence of Organic Chemistry Jane M. Cram, Donald J. Cram, 1978

Study Guide to Organic Chemistry, 4th Ed Robert Thornton Morrison, Robert Neilson

Boyd, 1983 A popular introduction to organic chemistry which stresses the importance of molecular structure in understanding the properties and principles of organic chemistry. Provides a wide variety of spectra to be analyzed. Features four color photographs throughout

**Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition** Susan McMurry, 2000 Provides answers and explanations to all in text and end of chapter exercises. Also includes

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## **Table of Contents Study Guide Organic Chemistry Answers**

1. Understanding the eBook Study Guide Organic Chemistry Answers
  - The Rise of Digital Reading Study Guide Organic Chemistry Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Organic Chemistry Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide Organic Chemistry Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Organic Chemistry Answers
  - Personalized Recommendations
  - Study Guide Organic Chemistry Answers User Reviews and Ratings
  - Study Guide Organic Chemistry Answers and Bestseller Lists
5. Accessing Study Guide Organic Chemistry Answers Free and Paid eBooks
  - Study Guide Organic Chemistry Answers Public Domain eBooks
  - Study Guide Organic Chemistry Answers eBook Subscription Services
  - Study Guide Organic Chemistry Answers Budget-Friendly Options
6. Navigating Study Guide Organic Chemistry Answers eBook Formats

- ePub, PDF, MOBI, and More
- Study Guide Organic Chemistry Answers Compatibility with Devices
- Study Guide Organic Chemistry Answers Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Study Guide Organic Chemistry Answers
- Highlighting and Note-Taking Study Guide Organic Chemistry Answers
- Interactive Elements Study Guide Organic Chemistry Answers

8. Staying Engaged with Study Guide Organic Chemistry Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Study Guide Organic Chemistry Answers

9. Balancing eBooks and Physical Books Study Guide Organic Chemistry Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Study Guide Organic Chemistry Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Study Guide Organic Chemistry Answers

- Setting Reading Goals Study Guide Organic Chemistry Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Study Guide Organic Chemistry Answers

- Fact-Checking eBook Content of Study Guide Organic Chemistry Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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