

PRACTICE TEST/THERMOCHEMISTRY/AP CHEMISTRY

1. a) Calculate the amount of heat transferred when 10.00 g of $\text{N}_2\text{O}(\text{g})$ is formed by the following reaction:



- b) Draw an energy diagram for this process.

2. Predict the value for ΔH°_f for the following scenarios:

- a) $\text{Br}_2(\text{g})$
- b) $\text{Br}_2(\text{l})$
- c) $\text{I}_2(\text{g})$
- d) $\text{I}_2(\text{s})$

3. Calculate the ΔH_{rxn} for the following reaction:



$$\begin{aligned}\Delta H^\circ_f \text{ C}_2\text{H}_4(\text{g}) &= 226.6 \text{ kJ/mol} \\ \Delta H^\circ_f \text{ CO}_2(\text{g}) &= -393.5 \text{ kJ/mol} \\ \Delta H^\circ_f \text{ H}_2\text{O}(\text{l}) &= -285.8 \text{ kJ/mol}\end{aligned}$$

4. A 5.00 g sample of liquid water at 25.0°C is heated by the addition of 84.0 J of energy. The final temperature of the water is _____ $^\circ\text{C}$. The specific heat capacity of the liquid is $4.18 \text{ J/g}^\circ\text{C}$.

- a) -21 b) 4.02 c) 25.2 d) 95.2 e) 29.0

5. *THIS PROBLEM WAS TAKEN DIRECTLY FROM THE 1995 AP CHEM. EXAM.*
Propane is a hydrocarbon that is commonly used as a fuel for cooking.
Propane's formula is C_3H_8 .

- a) Write a balanced equation for the complete combustion of propane gas.
- b) Calculate the volume of air at 30°C and 1.00 atm that is needed to burn completely 10.0 g of propane. Assume that air is 21.0% O_2 by volume.
- c) The heat of combustion ($\Delta H^\circ_{\text{combustion}}$) is $-2,220.1 \text{ kJ/mol}$. Calculate the heat of formation, ΔH°_f , of propane given that ΔH_f of $\text{H}_2\text{O}(\text{l})$ is -285.3 kJ/mol and ΔH°_f of $\text{CO}_2(\text{g})$ is -393.5 kJ/mol .
- d) Assuming that all of the heat evolved burning 10.0 g propane is transferred to 8.00 kg of water (specific heat = $4.184 \text{ J/g}^\circ\text{C}$), calculate the increase in temperature of the water.

Thermochemistry Practice Test

Albert S. Tarendash, Frederick J. Rowe



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