

PRACTICE TEST/THERMOCHEMISTRY/AP CHEMISTRY

1. a) Calculate the amount of heat transferred when 10.00 g of $\text{N}_2\text{O(g)}$ is formed by the following reaction:



b) Draw an energy diagram for this process.

2. Predict the value for ΔH°_f for the following scenarios:

- a) $\text{Br}_2\text{(g)}$
- b) $\text{Br}_2\text{(l)}$
- c) $\text{I}_2\text{(g)}$
- d) $\text{I}_2\text{(s)}$

3. Calculate the ΔH_{rxn} for the following reaction:



$$\Delta H^\circ_f \text{ C}_2\text{H}_4\text{(g)} = 236.6 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{ CO}_2\text{(g)} = -393.5 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{ H}_2\text{O(l)} = -285.8 \text{ kJ/mol}$$

4. A 5.00 g sample of liquid water at 25.0°C is heated by the addition of 84.0 J of energy. The final temperature of the water is _____ $^\circ\text{C}$. The specific heat capacity of the liquid is $4.18 \text{ J/g}^\circ\text{C}$.

- a) -21
- b) 4.02
- c) 25.2
- d) 95.2
- e) 29.0

5. THIS PROBLEM WAS TAKEN DIRECTLY FROM THE 1995 AP CHEM. EXAM.

Propane is a hydrocarbon that is commonly used as a fuel for cooking. Propane's formula is C_3H_8 .

- a) Write a balanced equation for the complete combustion of propane gas.
- b) Calculate the volume of air at 30°C and 1.00 atm that is needed to burn completely 10.0 g of propane. Assume that air is 21.0% O_2 by volume.
- c) The heat of combustion ($\Delta H^\circ_{\text{combustion}}$) is -2,220.1 kJ/mol. Calculate the heat of formation, ΔH°_f , of propane given that ΔH_f of $\text{H}_2\text{O(l)}$ is -285.3 kJ/mol and ΔH°_f of $\text{CO}_2\text{(g)}$ is -393.5 kJ/mol.
- d) Assuming that all of the heat evolved burning 10.0 g propane is transferred to 8.00 kg of water (specific heat = $4.184 \text{ J/g}^\circ\text{C}$), calculate the increase in temperature of the water.

Thermochemistry Practice Test

Albert S. Tarendash, Frederick J. Rowe

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