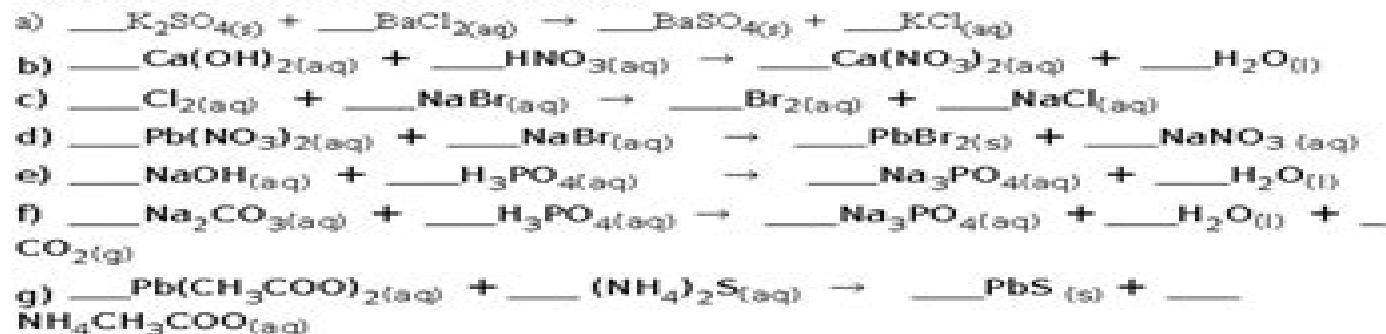


1. Balance the following equations. Then write NET ionic equations for each. Assuming 1L of total solution, calculate the concentration of each ion left in solution after reaction.



2. Calculate the number of:

(a) grams in 0.25 mol of magnesium hydroxide. (14.5 g)

(b) moles in 11.7 g of sodium chloride. (0.2 mol)

(c) moles in 4.9 g of hydrogen sulphate. (0.05 mol)

3. Calculate the molarity of solutions containing the following amounts of solute:

(a) 31.5 g HNO_3 per litre of solution. (0.5 M)

(b) 26.5 g Na_2CO_3 per 200 mL of solution. (1.25 M)

(c) 66.57 g AlCl_3 per 400 mL of solution. (1.25 M)

4. How many moles of solute are there in 100 mL of a 0.100 M solution? (0.01 mol)

5. How many grams of NaCl are there in 250 mL of a 2.50 M NaCl solution? (36.6 g)

6. What volume of a 0.75 M solution could be prepared from 500 g of Na_2SO_4 ? (4.7 L)

7. One litre of 12 M HCl is diluted to 20 L. What is the molarity of the

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Kevin Revell

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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshman in STEM fields.

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all the strengths of the previous editions while adding emphasis on conceptual understanding and critical thinking *The Software Encyclopedia 2000* Bowker Editorial Staff,2000-05 **Introduction to Chemistry, Laboratory Manual** T. R. Dickson,1994-12-23 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives Three themes are stressed throughout the text developing chemical thinking and a chemical vision learning problem solving methods and utilizing group work and discussion activities These themes involve and engage the students in their own learning processes they are challenged to be active The presentation of topics has been altered to include a new chapter which introduces the students to scientific thinking and shows that chemistry involves interesting and relevant topics The reorganization presents many core concepts in the first five chapters preparing students for later chapters In addition the author has added vignettes throughout the chapters referring to health technology the environment and society as well as to specific tools of direct use to students

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