

1. Balance the following equations. Then write **NET** ionic equations for each. Assuming 1L of total solution, calculate the concentration of each ion left in solution after reaction.



2. Calculate the number of:

- (a) grams in 0.25 mol of magnesium hydroxide. (14.5 g)  
 (b) moles in 11.7 g of sodium chloride. (0.2 mol)  
 (c) moles in 4.9 g of hydrogen sulphate. (0.05 mol)

3. Calculate the molarity of solutions containing the following amounts of solute:

- (a) 31.5 g  $\text{HNO}_3$  per litre of solution. (0.5 M)  
 (b) 26.5 g  $\text{Na}_2\text{CO}_3$  per 200 mL of solution. (1.25 M)  
 (c) 66.57 g  $\text{AlCl}_3$  per 400 mL of solution. (1.25 M)

4. How many moles of solute are there in 100 mL of a 0.100 M solution? (0.01 mol)  
 5. How many grams of NaCl are there in 250 mL of a 2.50 M NaCl solution? (36.6 g)  
 6. What volume of a 0.75 M solution could be prepared from 500 g of  $\text{Na}_2\text{SO}_4$ ? (4.7 L)  
 7. One litre of 12 M HCl is diluted to 20 L. What is the molarity of the

# Solution Stoichiometry Practice Problems

**Kevin Revell**



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