

### Unit 1 Study Guide — ANSWER KEY

1. Define scientific method: **a procedure for collecting, testing, and communicating information**
2. Define hypothesis: **a suggested answer to the question, an educated guess**
3. Define experiment: **a scientific test so see what will happen**
4. Define dependent variable: **the variable that you do NOT control. (on a graph is usually on the Y axis)**
5. Define independent variable: **the variable that I control. (on a graph is usually on the X axis)**
6. Define control: **the part of the experiment where everything is left the same.**
7. Define observation: **information collected with any of your 5 senses**
8. Define density: **the ratio of mass per unit volume (D = m/v)**
9. Define accuracy: **the measure of how close your data is to the ACTUAL or true value.**
10. Define precision: **a measure of how close your data (measurements) are to each other.**
11. Why do scientists use SI units of measurement? **So that measurements can be understood internationally (around the world).**
12. List (in the correct order) the steps of the scientific method.

- |                               |  |
|-------------------------------|--|
| 1. <b>make an observation</b> | 4. <b>test hypothesis (experiment)</b>   |
| 2. <b>ask a question</b>      | 5. <b>analyze data/form a conclusion</b> |
| 3. <b>form a hypothesis</b>   | 6. <b>develop a theory</b>               |

13. What mnemonic device is used to remember the metric system?

**K H Da B d c m**  
**g, m, L**

14. What is the sentence (mnemonic) that will help you remember the METRIC prefixes in order?

**King Henry Died By Drinking Chocolate Milk**

15. Fill in the table below describing the metric system:

	<b>Kilo</b>	<b>Hecto</b>	<b>Deca</b>	<b>Base</b>	<b>Deci</b>	<b>Centi</b>	<b>Milli</b>
Abbreviation	<b>K</b>	<b>H</b>	<b>Da</b>	<b>g, m, L</b>	<b>d</b>	<b>c</b>	<b>m</b>
Meaning in powers of 10	<b><math>10^3</math></b>	<b><math>10^2</math></b>	<b><math>10^1</math></b>	<b><math>10^0</math></b>	<b><math>10^{-1}</math></b>	<b><math>10^{-2}</math></b>	<b><math>10^{-3}</math></b>

#### Convert the following:

- |                                 |                                    |
|---------------------------------|------------------------------------|
| 16. 25L to <u>25000</u> mL      | 21. 25 cm to <u>0.25</u> m         |
| 17. 2.3 km to <u>230,000</u> cm | 22. 56784 mm to <u>0.056784</u> km |
| 18. 0.045 g to <u>45</u> mg     | 23. 15 g to <u>0.015</u> kg        |
| 19. 0.35 kg to <u>350</u> g     | 24. 3.452 kL to <u>3452</u> L      |
| 20. 300 mL to <u>0.300</u> L    | 25. 47 m to <u>47000</u> mm        |

#### Scientific Notation:

Convert the following to scientific notation:

- |   |   |
|---|---|
| 26) 0.005 <u><math>5 \times 10^{-3}</math></u>  | 30) 0.025 <u><math>2.5 \times 10^{-2}</math></u>  |
| 27) 5.050 <u><math>5.05 \times 10^1</math></u>  | 31) 0.0025 <u><math>2.5 \times 10^{-3}</math></u> |
| 28) 0.0008 <u><math>8 \times 10^{-4}</math></u> | 32) 500 <u><math>5 \times 10^2</math></u>         |
| 29) 1,000 <u><math>1 \times 10^3</math></u>     | 33) 5,000 <u><math>5 \times 10^3</math></u>       |

# Study Guide Unit Measurement Answers

**Albert Heuer, Louis Sinopoli**



## **Study Guide Unit Measurement Answers:**

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crucial This revision has several significant changes URL s are included at the ends of each chapter many of the Science News boxes are updated to show hot topics      **Study Guide for Goldfrank's Toxicologic Emergencies, 11th Edition**  
Rana Biary,Mary Ann Howland,Silas W. Smith,Richard S. Weisman,Neal A. Lewin,Lewis R. Goldfrank,Robert S.  
Hoffman,2022-08-22 The hands on study companion for the most respected rigorous medical toxicology text available  
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Kellogg,1923

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