

ANSWER KEY

ATOMS AND ELECTRONS STUDY GUIDE

Learning Goals: (a) (a.1-P.4.1) Explain a two-subshell theory of atomic orbitals (introduce atomic theory) by describing changes in the atomic model over time including those changes that are consistent with experimental evidence. (a.1-P.4.2) Explain the relationship between atomic orbitals (AOs) and orbitals (Orbitals) in terms of probability, nodal structures, radial probability plots and ψ^2 in terms of their nodes, electron charges, and locations within the atom. (a.1-P.4.3) Explain properties of atoms and their elements in the periodic table in terms of configurations of their electrons. (a.1-P.4.4) Explain the importance of orbitals in the four fundamental interactions (strong, electromagnetic, weak, gravity), using examples.

ANSWER KEY

Directions: Match the atomic theory selected on the correct card(s).

Model(s)	Model(s)	Atomic Theory (Select one or more)
1C	Continuous (a.1-P.4.1)	A. Everything is made of atoms (that cannot be created, divided or destroyed)
1d	Discrete (a.1-P.4.1)	B. Matter is made of atoms (atoms cannot be created, changed, made up, destroyed or created through nuclear transformation)
1e	Quantum (a.1-P.4.1)	C. "matter is indivisible (uncuttable)"
1f	Quantum (a.1-P.4.1)	D. Group Number = Valence Electrons
1g	Model (a.1-P.4.1)	E. Electrons Cloud
1h	Electrons (a.1-P.4.1)	F. Atoms are mostly empty space (the subatomic particles exist)

Directions: Use the periodic table to answer the following questions:

- 1C** = 1d, 1e. The number of protons equals the number of electrons.
- atomic mass minus the atomic number will equal the **Number of neutrons**.
- 1f, 1g** = **1h**. The number of protons and electrons (positively charged) are equal to the number of neutrons (negatively charged).

Study Guide Why Atoms Combine Answer Key

**Katherine J. Denniston, Denniston
Katherine**



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