

1.	What is the third law of thermodynamics?	The entropy of a perfect crystal is zero.
2.	What is the second law of thermodynamics?	As a whole, the universe always tends towards increasing entropy.
3.	What is the relationship between number of microstates and entropy?	More microstates = higher entropy
4.	What is the relationship between temperature and entropy?	Temperature increases, entropy increases (direct)
5.	What is the relationship between atomic weight and entropy?	Heavier = increased entropy (direct)
6.	What happens to entropy when a substance is dissolved in another?	Entropy increases
7.	What is the relationship between moles of gas and entropy?	When moles of gas increase, entropy increases (direct)
8.	What is the relationship between free energy, enthalpy, and entropy?	$\Delta G(\text{system}) = \Delta H(\text{system}) - T\Delta S(\text{system})$
9.	What does change in G measure?	The extent of the spontaneity of a process, and the useful energy available from it
10.	What is the relationship between work and spontaneous processes?	Spontaneous processes produce work
11.	What is the relationship between work and non-spontaneous processes?	Non-spontaneous processes require work

Study Guide Thermochemistry

Steven S. Zumdahl, Paul B. Kelter



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