

Phase Change Worksheet – Answer Sheet

- 1) A 12 oz. can of soda weighs about 450 grams. How many joules are released when a can of soda is cooled from 25 degrees Celsius (room temperature) to 4 degrees Celsius (the temperature of a refrigerator). The specific heat of liquid water is 4.184 J/g°C. **39.5 kJ**
- 2) How many joules are required to heat 250 grams of liquid water from 0 °C to 100 °C ? **104.5 kJ**
- 3) How many joules are required to melt 100 grams of water? The heat of fusion of water is 334 J/g. **33.4 kJ**
- 4) How many joules are required to heat 100 grams of water? The heat of vaporization of water is 2260 J/g. **226 kJ**
- 5) How many joules are required to heat 100 grams of water from 25 °C to 100 °C? The specific heat of water is 4.184 J/g°C. **104.5 kJ**
- 6) How many joules are given off when 100 grams of water are cooled from 25 °C to 100 °C? The specific heat of water is 4.184 J/g°C. **104.5 kJ**
- 7) How many joules are required to heat 10 grams of water from 25 °C to 100 °C? The specific heat of water is 4.184 J/g°C. **10.45 kJ**
- 8) How many joules are required to heat 100 grams of water from 25 °C to 100 °C? The heat of fusion of ice is 334 J/g. **33.4 kJ**
- 9) How many joules are released when 100 grams of liquid water are cooled from 100 °C to 25 °C? The heat of vaporization of water is 2260 J/g. **226 kJ**
- 10) How many joules are required to heat the temperature of 100 grams of water from 25 °C to 100 °C? The heat of fusion of water is 334 J/g. **33.4 kJ**

Study Guide Phase Changes Answer Key

Cheryl Hale, Crooks



Study Guide Phase Changes Answer Key:

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Resources for Teaching Middle School Science Smithsonian Institution, National Academy of Engineering, National Science Resources Center of the National Academy of Sciences, Institute of Medicine, 1998-04-30

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