

Chapter 19 Oxidation-Reduction Reactions

Chapter Opener

__ **Chapter Overview, TE** Review the objectives listed in the Student Edition.

Section 1 Oxidation and Reduction

PACING: 45 minutes

PENNSYLVANIA ACADEMIC STANDARDS FOR SCIENCE AND TECHNOLOGY:

3.1.E.1 Describe how fundamental science and technology concepts are used to solve practical problems.

Objectives

1. Assign oxidation numbers to reactant and product species.
2. Define *oxidation* and *reduction*.
3. Explain what an oxidation-reduction (redox reaction) is.

Focus (5 minutes)

__ **Lesson Starter, TE** Review oxidation numbers and have students familiarize themselves with the rules for assigning oxidation numbers in **Table I**.

MOTIVATE (10 minutes)

__ **Demonstration, A Spontaneous Redox Reaction, TE** Demonstrate a redox reaction by using KI, KMnO_4 , and HCl and make sure students can explain the formation of the precipitate and the color changes.

TEACH (20 minutes)

__ **PowerPoint**

__ **Visual Strategy, Figure 1, TE** Make sure students understand how this figure illustrates how an element can exist in a number of different oxidation states. (**GENERAL**)

__ **Chemistry in Action, Photochromic Lenses, SE** Photochromic eyeglasses darken and become transparent according to ultraviolet light because of an oxidation-reduction reaction.

__ **Application, TE** Sodium hypochlorite, NaOCl , is the chemical name for household bleach used to remove stains from clothing through a redox reaction.

__ **Discussion, TE** Students should follow up the spontaneous redox reaction demonstration by explaining the results in terms of oxidation and reduction and by writing formulas, using **Table I** for the oxidation numbers.

__ **Chemistry in Action, Skunk-Spray Remedy, SE** Paul Krebaum invented a better way than tomato juice to get rid of skunk spray odor.

Study Guide Redox Reactions Answers

Margaret G. Kimble



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general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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