

# Stoichiometry

## Section 11.1 Defining Stoichiometry

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### Practice Problems

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1. Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that the law of conservation of mass is observed.



1 molecule  $\text{N}_2$  + 3 molecules

$\text{H}_2 \rightarrow$  2 molecules  $\text{NH}_3$

1 mole  $\text{N}_2$  + 3 moles  $\text{H}_2 \rightarrow$  2 moles  $\text{NH}_3$

Mass:

$\text{N}_2$ : (2 mol)(14.007 g/mol) = 28.014 g

$3\text{H}_2$ : (6 mol)(1.008 g/mol) = 6.048 g

$2\text{NH}_3$ : (2 mol)(14.007 g/mol) +  
(6 mol)(1.008 g/mol) = 34.062 g

28.014 g  $\text{N}_2$  + 6.048 g  $\text{H}_2 \rightarrow$  34.062 g  $\text{NH}_3$

34.062 g reactants = 34.062 g products



1 molecule  $\text{HCl}$  + 1 formula unit  $\text{KOH} \rightarrow$

1 formula unit  $\text{KCl}$  + 1 molecule  $\text{H}_2\text{O}$

1 mole  $\text{HCl}$  + 1 mole  $\text{KOH} \rightarrow$

1 mole  $\text{KCl}$  + 1 mole  $\text{H}_2\text{O}$

Mass:

$\text{HCl}$ : (1 mol)(1.008 g/mol) +  
(1 mol)(35.453 g/mol) = 36.461 g

$\text{KOH}$ : (1 mol)(39.098 g/mol) +  
(1 mol)(15.999 g/mol) + (1 mol)(1.008 g/mol)  
= 56.105 g

$\text{KCl}$ : (1 mol)(39.098 g/mol) +  
(1 mol)(35.453 g/mol) = 74.551 g

$\text{H}_2\text{O}$ : (2 mol)(1.008 g/mol) +  
(1 mol)(15.999 g/mol) = 18.015 g

36.461 g  $\text{HCl}$  + 56.105 g  $\text{KOH} \rightarrow$   
74.551 g  $\text{KCl}$  + 18.015 g  $\text{H}_2\text{O}$

92.566 g reactants = 92.566 g products



2 atoms  $\text{Mg}$  + 1 molecule

$\text{O}_2 \rightarrow$  2 formula units  $\text{MgO}$

2 moles  $\text{Mg}$  + 1 mole  $\text{O}_2 \rightarrow$  2 moles  $\text{MgO}$

Mass:

$2\text{Mg}$ : (2 mol)(24.305 g/mol) = 48.610 g

$\text{O}_2$ : (2 mol)(15.999 g/mol) = 31.998 g

$2\text{MgO}$ : (2 mol)(24.305 g/mol) +  
(2 mol)(15.999 g/mol) = 80.608 g

48.610 g  $\text{Mg}$  + 31.998 g  $\text{O}_2 \rightarrow$  80.608 g  $\text{MgO}$

80.608 g reactants = 80.608 g products

2. **Challenge** For each of the following, balance the chemical equation; interpret the equation in terms of particles, moles, and mass; and show that the law of conservation of mass is observed.



2 atoms  $\text{Na}$  + 2 molecules  $\text{H}_2\text{O} \rightarrow$

2 formula units  $\text{NaOH}$  + 1 molecule  $\text{H}_2$

2 mol  $\text{Na}$  + 2 mol  $\text{H}_2\text{O} \rightarrow$

2 mol  $\text{NaOH}$  + 1 mol  $\text{H}_2$

Mass:

$2\text{Na}$ : (2 mol)(22.990 g/mol) = 45.980 g

$2\text{H}_2\text{O}$ : (4 mol)(1.008 g/mol) +  
(2 mol)(15.999 g/mol) = 36.030 g

$2\text{NaOH}$ : (2 mol)(22.990 g/mol) +  
(2 mol)(15.999 g/mol) + (2 mol)(1.008 g/mol)  
= 79.994 g

$\text{H}_2$ : (2 mol)(1.008 g/mol) = 2.016 g

45.980 g  $\text{Na}$  + 36.030 g  $\text{H}_2\text{O} \rightarrow$

79.994 g  $\text{NaOH}$  + 2.016 g  $\text{H}_2$

82.01 g reactants = 82.01 g products

# Study Guide Stoichiometry Answer Key

**Steven S. Zumdahl, Paul B. Kelter**



## **Study Guide Stoichiometry Answer Key:**

*Study Guide to Accompany Basics for Chemistry* Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

Chemistry, Student Study Guide John A. Olmsted, Gregory M. Williams, 2002-01-04 This third edition continues to innovate by providing students with an integrated and modern approach to the subject The text emphasizes the modern tools of chemistry while incorporating historical evidence and its unique molecular quantitative emphasis is further reinforced by an integrated media package developed by the authors Also of benefit is the just in time presentation of key content only providing details once they are needed While key topics and analytical techniques have been updated there is now an additional third chapter on chemical equilibrium The authors have also developed an expanded and more integrated problem solving emphasis that now incorporates a 4 step strategy throughout complete with text icons The whole is backed by a range of supplements including a new illustration program a tutorial CD interactive learningware an extensive Web CT component an instructor's resource CD and a solution CD

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life

on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement Class 11-12 Chemistry MCQ (Multiple Choice Questions) Arshad Iqbal, 2019-05-17 The Class 11 12 Chemistry Multiple Choice Questions MCQ Quiz with Answers PDF College Chemistry MCQ PDF Download Quiz Questions Chapter 1 6 Practice Tests with Answer Key 11th 12th Grade Chemistry Questions Bank MCQs Notes includes revision guide for problem solving with hundreds of solved MCQs Class 11 12 Chemistry MCQ with Answers PDF book covers basic concepts analytical and practical assessment tests Class 11 12 Chemistry MCQ PDF book helps to practice test questions from exam prep notes The Class 11 12 Chemistry MCQs with Answers PDF eBook includes revision guide with verbal quantitative and analytical past papers solved MCQs Class 11 12 Chemistry Multiple Choice Questions and Answers MCQs PDF Free download chapter 1 a book covers solved quiz questions and answers on chapters atomic structure basic chemistry chemical bonding chemistry experimental techniques gases liquids and solids tests for college and university revision guide Class 11 12 Chemistry Quiz Questions and Answers PDF free download eBook s sample covers beginner s solved questions textbook s study notes to practice online tests The book Grade 11 12 Chemistry MCQs Chapter 1 6 PDF includes college question papers to review practice tests for exams Class 11 12 Chemistry Multiple Choice Questions MCQ with Answers PDF digital edition eBook a study guide with textbook chapters tests for NEET MCAT GRE GMAT SAT ACT competitive exam College Chemistry Mock Tests Chapter 1 6 eBook covers problem solving exam tests from chemistry textbook and practical eBook chapter wise as Chapter 1 Atomic Structure MCQ Chapter 2 Basic Chemistry MCQ Chapter 3 Chemical Bonding MCQ Chapter 4 Experimental Techniques MCQ Chapter 5 Gases MCQ Chapter 6 Liquids and Solids MCQ The Atomic Structure MCQ PDF e Book Chapter 1 practice test to solve MCQ questions on Atoms atomic spectrum atomic absorption spectrum atomic emission spectrum molecules azimuthal quantum number Bohr s model Bohr s atomic model defects charge to mass ratio of electron discovery of electron discovery of neutron discovery of proton dual nature of matter electron charge electron distribution electron radius and energy derivation electron velocity electronic configuration of elements energy of revolving electron fundamental particles Heisenberg s uncertainty principle hydrogen spectrum magnetic quantum number mass of electron metallic crystals properties Moseley law neutron properties orbital concept photons wave number Planck s quantum theory properties of cathode rays properties of positive rays quantum numbers quantum theory Rutherford model of atom shapes of orbitals spin quantum number what is spectrum x rays and atomic number The Basic Chemistry MCQ PDF e Book Chapter 2 practice test to solve MCQ questions on Basic chemistry atomic mass atoms molecules Avogadro s law combustion analysis empirical formula isotopes mass spectrometer molar volume molecular ions moles positive and negative ions relative abundance spectrometer and stoichiometry The Chemical Bonding MCQ PDF e Book Chapter 3 practice test to solve MCQ questions on Chemical bonding chemical combinations atomic radii atomic radius periodic table atomic ionic and covalent radii atoms and

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**Study Guide, Chemical Principles, Fifth Edition, Zumdahl** Paul B. Kelter, Steven S. Zumdahl, 2004-04

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**Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **Holt Chemistry** Holt Rinehart & Winston, 2003-01-24 *Essentials of Introductory Chemistry* Russo Steve Silver Michael, Steve Russo, 2001-12

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