

The Mole Study Guide

Mass Composition: What's your element's %? given after the decimal? Molar mass to obtain your answer!

10. There are 3.75 moles of CO_2 .

11. There are _____ moles of _____.

12. There are _____ moles of _____.

13. There are _____ moles of _____.

14. There are 2.5 moles of _____.

15. There are _____ moles of _____.

16. There are _____ moles of _____.

17. There are _____ moles of _____.

18. There are 4.5 moles of _____.

19. There are _____ moles of _____.

20. There are _____ moles of _____.

21. There are 3.5 moles of _____.

22. There are _____ moles of _____.

23. There are _____ moles of _____.

24. There are 2.5 moles of _____.

25. There are _____ moles of _____.

26. There are _____ moles of _____.

Name _____ Period _____ Date _____

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Fill in the Blank boxes for the image below.



Calculate the Molar Mass for the following Compounds.

10. H_2O

21. H_2O

Calculate the Percent Composition for each element in the following Compound.

10. H_2O

21. H_2O

22. H_2O

23. H_2O

24. H_2O

25. H_2O

26. H_2O

27. H_2O

28. H_2O

29. H_2O

30. H_2O

Write the Empirical Formula for each compound.
(Use subscripts to show the relative ratio of atoms.)

Molecular	Empirical
$\text{C}_6\text{H}_{12}\text{O}_6$	
$\text{C}_4\text{H}_8\text{O}_4$	
$\text{C}_2\text{H}_4\text{O}_2$	
$\text{C}_3\text{H}_6\text{O}_3$	
$\text{C}_5\text{H}_{10}\text{O}_5$	

Determine the Empirical and Molecular Formula for the compound given the information below.

10. A compound has the following percent composition: 48.00% Hg, 27.60% Cl, 24.40% N. The molar mass for Astatine is 210.0 g/mol.

a. If you have 100 g of Astatine, how many grams of Hg, Cl, and N do you have?

b. Determine the Empirical Formula for Astatine.

c. Determine the Molecular Formula for Astatine.

Study Guide The Mole

Stanley Johnson



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