

Name: _____

Date: _____

Subatomic Particles for Atoms, Ions and Isotopes-Practice

Use your periodic table to complete this worksheet.

	Atom/Ion/Isotope name	Standard Atomic Notation	Atomic Mass	Atomic Number	Number of Protons	Number of Electrons	Charge	Number of Neutrons
1	Strontium (isotope)	$^{90}_{38}\text{Sr}$	90	38	38	38	0	52
2	Beryllium (ion)	$^{9}_{4}\text{Be}^{2+}$	9	4	4	2	+2	5
3	Selenium (ion)	$^{79}_{34}\text{Se}^{2-}$	79	34	34	36	-2	45
4	Aluminum (ion)	$^{27}_{13}\text{Al}^{3+}$	27	13	13	10	+3	14
5	Helium (isotope)	$^{3}_{2}\text{He}$	3	2	2	2	0	1
6	Arsenic (atom)	$^{75}_{33}\text{As}$	75	33	33	33	0	42
7	Lithium (isotope)	$^{11}_{3}\text{Li}$	11	3	3	3	0	8
8	Sodium (ion)	$^{23}_{11}\text{Na}^{1+}$	23	11	11	10	+1	12
9	Mercury (atom)	$^{201}_{80}\text{Hg}$	201	80	80	80	0	121
10	Iron (isotope)	$^{53}_{26}\text{Fe}$	53	26	26	26	0	27
11	Nitrogen (ion)	$^{14}_{7}\text{N}^{3-}$	14	7	7	10	-3	7
12	Nickel (ion)	$^{59}_{28}\text{Ni}^{3+}$	59	28	28	25	+3	31
13	Chlorine (isotope)	$^{36}_{17}\text{Cl}$	36	17	17	17	0	19
14	Xenon (atom)	$^{131}_{54}\text{Xe}$	131	54	54	54	0	77
15	Zinc (isotope)	$^{59}_{30}\text{Zn}$	59	30	30	30	0	29

*NOTE: when you are finished, you should have 3 atoms, 6 ions and 6 isotopes.

Subatomic Particle Gizmo Answer Sheet

Randy L. Bell

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Table of Contents Subatomic Particle Gizmo Answer Sheet

1. Understanding the eBook Subatomic Particle Gizmo Answer Sheet
 - The Rise of Digital Reading Subatomic Particle Gizmo Answer Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Subatomic Particle Gizmo Answer Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Subatomic Particle Gizmo Answer Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Subatomic Particle Gizmo Answer Sheet
 - Personalized Recommendations
 - Subatomic Particle Gizmo Answer Sheet User Reviews and Ratings
 - Subatomic Particle Gizmo Answer Sheet and Bestseller Lists
5. Accessing Subatomic Particle Gizmo Answer Sheet Free and Paid eBooks
 - Subatomic Particle Gizmo Answer Sheet Public Domain eBooks
 - Subatomic Particle Gizmo Answer Sheet eBook Subscription Services
 - Subatomic Particle Gizmo Answer Sheet Budget-Friendly Options
6. Navigating Subatomic Particle Gizmo Answer Sheet eBook Formats

- ePub, PDF, MOBI, and More
- Subatomic Particle Gizmo Answer Sheet Compatibility with Devices
- Subatomic Particle Gizmo Answer Sheet Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Subatomic Particle Gizmo Answer Sheet
- Highlighting and Note-Taking Subatomic Particle Gizmo Answer Sheet
- Interactive Elements Subatomic Particle Gizmo Answer Sheet

8. Staying Engaged with Subatomic Particle Gizmo Answer Sheet

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Subatomic Particle Gizmo Answer Sheet

9. Balancing eBooks and Physical Books Subatomic Particle Gizmo Answer Sheet

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Subatomic Particle Gizmo Answer Sheet

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Subatomic Particle Gizmo Answer Sheet

- Setting Reading Goals Subatomic Particle Gizmo Answer Sheet
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Subatomic Particle Gizmo Answer Sheet

- Fact-Checking eBook Content of Subatomic Particle Gizmo Answer Sheet
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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