

Subatomic Particles for Atoms, Ions and Isotopes-Practice

Use your periodic table to complete this worksheet.

	Atom/Ion/Isotope name	Standard Atomic Notation	Atomic Mass	Atomic Number	Number of Protons	Number of Electrons	Charge	Number of Neutrons
1	Strontium (isotope)	$^{90}_{38}\text{Sr}$	90	38	38	38	0	52
2	Beryllium (ion)	$^9_4\text{Be}^{2+}$	9	4	4	2	+2	5
3	Selenium (ion)	$^{79}_{34}\text{Se}^{2-}$	79	34	34	36	-2	45
4	Aluminum (ion)	$^{27}_{13}\text{Al}^{3+}$	27	13	13	10	+3	14
5	Helium (isotope)	^3_2He	3	2	2	2	0	1
6	Arsenic (atom)	$^{75}_{33}\text{As}$	75	33	33	33	0	42
7	Lithium (isotope)	$^{11}_3\text{Li}$	11	3	3	3	0	8
8	Sodium (ion)	$^{23}_{11}\text{Na}^{1+}$	23	11	11	10	+1	12
9	Mercury (atom)	$^{201}_{80}\text{Hg}$	201	80	80	80	0	121
10	Iron (isotope)	$^{53}_{26}\text{Fe}$	53	26	26	26	0	27
11	Nitrogen (ion)	$^{14}_7\text{N}^{3-}$	14	7	7	10	-3	7
12	Nickel (ion)	$^{59}_{28}\text{Ni}^{3+}$	59	28	28	25	+3	31
13	Chlorine (isotope)	$^{36}_{17}\text{Cl}$	36	17	17	17	0	19
14	Xenon (atom)	$^{131}_{54}\text{Xe}$	131	54	54	54	0	77
15	Zinc (isotope)	$^{59}_{30}\text{Zn}$	59	30	30	30	0	29

*NOTE: when you are finished, you should have 3 atoms, 6 ions and 6 isotopes.

Subatomic Particle Gizmo Answer Sheet

Randy L. Bell

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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